

## MDCAT Chemistry Chapter 9 Electrochemistry Online Test

Sr	Questions	Answers Choice
1	Octet rule is not allowed in the formation of	A. NF <sub>3</sub> B. B.CF <sub>4</sub> C. CCl <sub>4</sub> D. <b>PCl<sub>5</sub></b>
2	what is the exact value of angle in BF <sub>3</sub>	A. 90 B. 104.51 C. 119.5 D. <b>120°</b>
3	pi-bond can be formed by sideways overlap of	A. s-orbital B. d-orbital C. <b>p-orbital</b> D. sp orbital
4	Total number of valence electrons in phosphonium ion (PH <sub>4</sub> <sup>+</sup> ) is	A. <b>8</b> B. 9 C. 12 D. 10
5	Elements of group IA and IIA are	A. electronegative B. neutral C. <b>electropositive</b> D. non-metals
6	Ionic bond is produced after complete transfer of	A. nucleus B. neutrons C. <b>electrons</b> D. protons
7	Elements having high ionization potential values are	A. metals B. <b>non- metal</b> C. liquids D. solid
8	Greater shielding effect corresponds to ionization potential value	A. greater B. <b>lesser</b> C. remain same D. no effect
9	Energy required to remove electron from an atom	A. <b>Ionization potential</b> B. Electronegativity C. Electropositivity D. <div>Electron affinity</div>
10	Covalent bonds are	A. <b>directional</b> B. Bidirectional C. Multidirectional D. Non directional
11	Bonding in MgO is an example of	A. <b>Ionic bond</b> B. Polar bond C. Covalent bond D. Coordination covalent bond
12	A covalent bond may be	A. 100% covalent B. <b>Partial ionic</b> C. 100% ionic D. Both a and b
13	Polarity of a molecule is expressed in terms of	A. Bond strength B. <b>Dipole moment</b> C. Bond length D. Shape
14	Carbon-Carbon double bond length in C <sub>3</sub> H <sub>6</sub>	A. 154 pm B. <b>134 pm</b> C. 120 pm D. 105 pm
15	Geometry of simple molecule with sp <sup>2</sup> hybridization	A. <b>Triangular planar</b> B. Trigonal C. Square planner D. Pyramidal

16	Which one of the followings has polar covalent bonds hut is overall nom-polar molecule:	A. HF B. $\text{CO}_2$ C. $\text{CH}_4$ D. $\text{N}_2$
17	Mostly ionic compounds are produced between elements of group	A. IA and IIA B. IB and VIB C. IA, IIA and VII-A D. IA and IB
18	Bond will be covalent when electronegativity difference of bonded atom is	A. Equal to 1.7 B. between 0.5 to 1.7 C. Greater to 1.7 D. zero
19	Energy of atom in compound is	A. Higher than individual atom B. Lower than individual atom C. equal to individual atom D. Impossible to predict