MDCAT Chemistry Chapter 8 Thermo-chemistry and Energetics of chemical reactions Online Test

Sr	Questions	Answers Choice
1	On ascending the electrochemical series strength as reducing agent	A. Increases B. Decreases C. Remains same D. not determinable
2	The potential difference set up at 25 C and 1 atm when clectrode is dipped m Tis one molar ionic sohution is called	A. Single electrode potential B. electrode potential C. Standard electrode potential D. Standard hydrogen electrode
3	The cell which converts electrical energy to chemical energy is called	A. Electrochemical cell B. Voltaic cell C. Galvanic cell D. Down's cell
4	Molten lead and iead (II) bromide both conduct electricity. Which one of the following statements relating to this is true?	A. Both undergo chemical change when they conduct B. Both conduct by the movement of charge particles C. Both will also conduct in the solid state D. Both contain mobile electrons
5	Zinc reacts with dilute acids to liberate hydrogen. This is because:	A. Zn2+ ion is a powertul osidising agent than H' ion B. H+ ion is a powerful oxidising agent than Zn ion C. Zn2+ ion is a powerful reducing agent than H' ion D. H+' ion is a powerful reducing agent than Zn- ian
6	Which of the following salts would give the same products irrespective of whether its molten form or concentrated aqueous solution is electrolysed?	A. Magnesium bromide B. Magnesium sulphate C. Copper sulphate D. Copper chloride
7	Electrolytic products of dilute aqueous solution of sodium sulphate is	A. Na. SO2 B. H2. S02 C. Na. O2 D. H2,O2
8	Which of the following is an application of electrochemical series	A. Prediction of the feasibility of chemical reaction B. Calculation of the cell voltage C. Prediction of reaction of metal with dilute acid D. All of the above
9	If a strip of Cu metal is placed in a solution of FeSO4	A. Cu will be deposited B. Cu and Fe both dissolve C. Fe is precipitated out D. No reaction take place
10	Which of the following statements is not correct about galvanic cell?	A. Anode is negatively charged B. Cathode is positively charged C. Reduction occurs at anode D. Reduction occurs at cathode
11	By the electrolysis of CuCl2 using inert electrodes of platinum which species is deposited at cathode	A. H2 B. O2 C. Cu D. Cl
12	The reduction potentials of non-metals are A =+0.54V, B=+1.08V, C=+1.36V. D= +2.87V Which non -metal can displace all other from aqucous solution of their salts	A. A B. C C. B D. D
13	Which of following is oxidation state of oxygen in peroxides?	A. -2 B. 1/2 C. -1 D. +2

14	Stronger the oxidizing agent, higher is	A. Redox potential B. Standard reduction potential C. Reduction potential D. _{Oxidation potemial}
15	The emf produced by galvanic cell is called	A. Cell potential B. Oxidation potential C. Redox potential D. Reduction potential
16	If Cl2 is passed through hot NaOH. NaClO3 is formed and the oxidation number of Cl changes from	A. -1 to 0 B. 0 to +5 C. 0 to-1 D. 0 to +1
17	The element which has greatest value of Reduction potential is used as	A. Strongest reducing agent B. Weak oxidizing and strong reducing agent C. Strongest oxidizing agent D. None of these