

MDCAT Chemistry Chapter 8 Thermo-chemistry and Energetics of chemical reactions Online Test

Sr	Questions	Answers Choice
1	The oxidation state of carbon in $\text{C}_2\text{O}^{2-}_4$ is	A. +4 B. -4 C. +3 D. +2
2	The common oxidation number of halogens is	A. -1 B. +1 C. -2 D. 0
3	In SO^{2-}_4 the oxidation number of sulphur is	A. -8 B. -6 C. +8 D. +6
4	In MgCl_2 , the oxidation state of Cl is	A. Zero B. -2 C. +2 D. -1
5	In all oxidation reactions, atoms of an element in a chemical species lose electrons and increases their	A. Oxidation states B. Reduction states C. Electrode D. Negative charges
6	The reaction which is responsible for the production of electricity in the voltaic cell is	A. Hydrolysis B. Oxidation C. Reduction D. Redox
7	The element with highest E°_{red}	A. N B. F C. O D. Cl
8	If a salt bridge is removed from two half cells the emf is	A. Increased B. Decreased C. Dropped to zero D. Electrodes will be reversed
9	SHE acts as anode when connected with Cu electrode but act as cathode with Zn electrode	A. Zn has less reduction potential than hydrogen and Cu B. Zn has high reduction potential than hydrogen and Cu C. Zn is below electrochemical series than hydrogen and Cu D. Zn has least tendency to lose electron
10	The electrochemical series is based on	A. pH scale B. Redox scale C. Hydrogen scale D. Arrhenius scale
11	The potential of SHE is taken as zero which is a value	A. Reference B. Arbitrary C. Exact D. Experimental
12	The working condition/s for SHE	A. 1atm pressure B. 1M H ⁺ solution C. 298K temperature D. All of these
13	The electrochemical reactions occurring at both the electrodes along with the electrolytic conduction constitute	A. Oxidation B. reduction C. Redox reaction D. electrolysis
		A. Na and H ₂

14	During the electrolysis of Fused NaCl, the products are	B. Na and Cl ₂ C. Na and O ₂ D. H ₂ and Cl ₂
15	The products of electrolysis of which of the followings are known	A. Fused electrolyte B. Aqueous solution of electrolyte C. Solid electrolyte D. Solid metal
16	Coinage metals Cu, Ag, and Au are the least reactive because they have	A. Negative reduction potential B. Positive reduction potential C. Negative oxidation potential D. Positive oxidation potential
17	Only those metals can replace Hydrogen from dilute acids, which have	A. High negative reduction potential B. Low negative reduction potential C. High positive reduction potential D. Low positive reduction potential
18	Which one of the following metals can replace the Copper from aqueous solution of its salt more easily?	A. Cd B. Fe C. Zn D. Na
19	The standard reduction potential of Zinc is	A. 0.76V B. 0.34 C. -0.34V D. -0.76V
20	When a metal rod is dipped in its one molar ionic solution	A. Electricity is produced B. Electricity is consumed C. Redox reaction occurs D. Potential difference is set up