MDCAT Chemistry Chapter 7 Reaction Kinetics Online Test

Sr	Questions	Answers Choice
1	The enthalpy of formation of a compound is	A. Positive B. Either positive or negative C. Negative D. None
2	The values of ΔH for the process I(g)+e-1---------> I-1(g) is:	A. >0 B. <0 C. 0 D. None
3	Calorie is equivalent to	A. 0.4184J B. 4.184J C. 418.4J D. 40.18J
4	If a reaction involves only solids and liquids, which of the following is true?	A. ΔH = ΔE B. ΔH =ΔE C. ΔH>ΔE D. AH =AE +nRT
5	NaOH+HCl- NaCl+ H2O. Enthalpy change in the above reaction is called	A. Enthalpy of reaction B. Enthalpy of Neutralisation C. Enthalpy of formation D. Enthalpy of combustion
6	Hess's law is analogous to	A. Law of heat summation B. law of increasing entropy C. Law of heat exchange D. Ist law of thermodynamics
7	The net heat change in a chemical reaction is the same whether it is brought about in two or more different ways in one or several steps.it is known as	A. Henry's law B. Hess's law C. joule's law D. Law of conservation of energy
8	Which of the following has positive value of enthalpy	A. Neutralisation B. Atomization C. combustion D. All of the above
9	For an endothermic reaction, enthalpy of reactants	A. Is smaller than that of the products B. Is greater than that of the products C. Must be greater or smaller than that of the products D. Is equal to that of the products
10	Enthalpy of neutralization of strong acids and strong bases have same values because	A. Neutralization leads to the formation of salt and water B. Acids always give rise to H+ and bases always furnish OH- C. Strong acids and bases are ionic substances D. The net change involves the combination of H and OH ions to form water