

MDCAT Chemistry Chapter 7 Reaction Kinetics Online Test

Sr	Questions	Answers Choice
1	Change in enthalpy (ΔH) of a system can be calculated by	A. $\Delta H = \Delta E - PV$ B. $\Delta H = \Delta E + q$ C. $\Delta H = \Delta E - q$ D. $\Delta H = \Delta E + P\Delta V$
2	Choose from the followings the correct statement about Born Haber cycle	A. Born Haber cycle is different from Hess's law B. The energy changes in a cyclic process is not zero C. The lattice energy of crystalline substances can be calculated easily D. None
3	Enthalpy of formation of one mole of ionic compound from gaseous ion under standard condition is called	A. Gibb's energy B. Gibb's energy C. Bond energy D. Lattice energy
4	According to Hess's law, the enthalpy change for a reaction	A. Depends on path B. Independent of the path C. The sum of ΔE and ΔH D. None of these
5	Decomposition of H_2O is	A. Endothermic reaction B. Nuclear reaction C. Exothermic reaction D. Zero nuclear reaction
6	The lattice energy of $NaCl$ is	A. 787 J/ mole B. 790 kJ/mol C. 780 kJ/ mol D. -787 kJ / mole
7	The value of ΔV being very small. The term $P\Delta V$ can be neglected for process involving	A. Liquid and gas B. Solids and gases C. Liquid and solid D. None of these
8	The enthalpy change for the reaction $C_2H_2 + 5/2 O_2 \rightarrow 2CO_2 + H_2O$ is known as enthalpy of	A. Formation of CO_2 B. Fusion of C_2H_4 C. Combustion of C_2H_4 D. Vaporization of C_2H_2
9	A state function which describes together the internal energy and product of pressure and volume is called	A. Enthalpy B. Internal energy C. Work D. Kinetic energy
10	The enthalpies of all elements in their standard states are	A. Unity B. always +ve C. always -ve D. zero
11	Total heat content of a system is called	A. Internal energy B. Entropy C. Enthalpy D. All of these
12	The measurement of enthalpy change at standard conditions means that we should manage the measurement at	A. $24^\circ C$ at 1 atm B. $25^\circ C$ at 1 atm C. $0^\circ C$ at 1 atm D. $100^\circ C$ 1 atm
13	Most of the reactions which give stable products are	A. Endothermic B. Exothermic C. Isothermal D. Non of these
14	During an exothermic or endothermic reaction which one of the following formula is used to calculate the amount of heat evolved or absorbed	A. $\Delta H = \Delta E + PV$ B. $\Delta E = q + w$ C. $\Delta p = \Delta H$ D. $q = m \times s \times \Delta T$
		A. Evaporation

15	The exothermic process is	B. Sublimation C. Respiration D. Boiling
16	The heat of reaction depends upon	A. Temperature of the reactants B. Physical states of the reactants and the products C. Both A) and B) D. Path of the reaction and the temperature
17	One Joule is equivalent to	A. 4.184 cal. B. 0.4184cal. C. 1/2 cal. D. 1/4.184 cal
18	If an endothermic reaction is allowed to take place very rapidly in air, the temperature of the surrounding air will	A. Remains constant B. Increase C. Decrease D. Either increase or decrease E. One Joule is equivalent to
19	What is not correct about ΔH_f	A. It is always negative B. Its value gives an idea about the relative stability of reactants and the products. C. Its value can be greater or less than zero D. Value depends upon nature of bonds
20	What is correct about heat of combustion	A. It is applicable to gaseous substances only B. It is always negative C. It is always positive D. It is positive in some cases while negative in other