MDCAT Chemistry Chapter 5 Solids Online Test

Sr	Questions	Answers Choice
1	Plastics are amorphous solids and	A. have sharp melting points B. undergo clean cleavage when cut with knife C. do not undergo clean cleavage D. possess orderly arrangement over long distances
2	Amorphous means	A. arranged B. ordered C. shaped D. shapeles (no arrangements)
3	The arrangement ABC, ABC .... is referred as	A. cubic close packing B. octahedral close packing C. hexagonal close packing D. tetrahedral close packing
4	All the metal shine when they are freshly cut The reason is	A. the conductivity of the metal is increased B. the process of cutting gives energy to the metal atoms C. the electrons become less delocalized according to valance bond theory D. the electrons are excited at higher energy levels and emit the photons when they fall back
5	The electrical conductivity of the metals decreases with the increasing temperature. This is because	A. the number of free electrons decrease B. the bonds of the metal atoms become weak C. the to and fro motion of the metal ions decrease D. the increase of to and fro motion of the metal ions hinders the free movement of electrons
6	Metallic bonds have been explained by many theories. Luis Pauling has proposed a theory called	A. molecular orbital theory B. electron gas theory C. band theory D. valence bond theory
7	How temperature affects the electrical conductivity of metals?	A. Does not change at all B. Decreases with increasing temperature C. Increases with increasing temperature D. Decreases with decreasing temperature
8	Which attractive forces cause molecular solids to be formed?	A. Ionic B. Metallic C. Covalent D. van der Waals
9	in diamond a unit cell is tetrahedral and averall crystai structure is	A. face centred cubic B. body centred cubic C. tetrahedral D. hexagonal
10	In diamond, which hybridization is there?	A. sp2 B. dsp2 C. sp3 D. sp
11	Which of the following solids does not have a covalent bond?	A. Silica B. Copper C. Diamond D. Graphite
12	The geometry of diamonds is	A. tetragonal B. cubic C. rhombohedral D. none of these

13	The molecules of CO_2, in dry ice form the	A. covalent crystals B. molecular crystals C. none of these crystals D. ionic crystals
14	The nature of crystal of diamond is	A. metallic B. molecular C. covalent D. ionic
15	The crystals formed due to London forces of interaction are	A. ionic B. covalent C. molecular D. metallic
16	In most of the cases the molecular crystals are	A. very soft B. soft C. extremely hard D. sufficiently hard
17	The crystal of diamond is	A. ionic B. molecular C. covalent D. metallic
18	Substance that does not show the process of sublimation is	A. $K_2Cr_2O_7$ B. iodine C. naphthalene D. NH_4Cl
19	The number of Na^+, ions which surround each Cl^- ion in the NaCl crystal lattice is	A. 8 B. 12 C. 6 D. 10
20	NaCl has face centered cubic structure. The Na ion at the face of the unit cell is shared by	A. 2-unit cells B. 4-units cells C. only one unit cell D. 8-unit cells