MDCAT Chemistry Chapter 1 Introduction to fundamental concepts of chemistry Online Test

Sr	Questions	Answers Choice
1	Glycerine is a polar compound. It boils at 290°C under one atmospheric pressure. It should be distilled under reduced pressure due to reason that	A. there are strong intermolecular forces between molecules of glycerine B. it decomposes at 290°C C. low pressure makes the liquid to boil at high temperature D. the reduced pressure decreases the boiing point of liquids
2	To cook the food at a high mountain is difficult as compared to at sea level. The reason is that:	A. the temperature at the top of the mountain is low B. the density of water decreases at the mountains C. the boiling point of water decreases at the mountain D. the hydrogen bonding in water changes with the change of height
3	CO2 and SO2 are both triatomic molecules, but heat of vaporization of SO2 is greater than that of CO2. This is due to	A. greater electronegative character of sulphur B. greater size of SO2 molecule C. SO2 is polar and CO2 is non-polar D. SO2 is more acidic in nature than CO2
4	The value of the vapour pressure of water at its boiling point at Karachi and Murree is	A. same B. different C. depends upon the environmental conditions in both cities D. greater at Murree and less at Karachi
5	Evaporation occurs at all temperatures and is effected by	A. surface area B. temperature C. intermolecular forces D. all of these
6	What s the boling point of H2O at the peak of Mount Everest?	A. 101 C° B. 69°C C. 100 C° D. 98° C
7	The hydrocarbon with maximum B.P is	A. CH4 B. C6H14 C. C4H10 D. C2H6
8	Point out that which is not an application of liquid crystals?	A. Source of energy B. In display of electrical devices C. For skin thermography D. As temperature sensor
9	Which of the following is not the property of liquid crystal	A. anisotropic B. isotropic C. three dimensional arrangement D. fluidity
10	Cholestryl benzoate tums into milky liquid at	A. 140°C B. 145°C C. 148C° D. 149°C
11	Water may boil at 120 °C when external pressure is:	A. 100 mm of Hg B. 700 mm of Hg C. 760 mm of Hg D. 1489 mm of Hg
12	The B.P of H2O at Murree Hills is	A. 99.8C B. 98°C C. 100C° D. 89°C
13	Styrene has empirical formula CH, and there is 92.2%C and 7.75% hydrogen. If molar mass is 104g mol^{-}, what will be integral multiple (n) to get molecular formula:	A. 2 B. 4 C. 6

		D. 8
14	The sole produets of combustion analysis are	A. CO2 and NH3 B. H2O and Mg(ClO4)2 C. CO2 and KOH D. CO2 and H2O
15	6Na+ Fe2O3 -------- 3 Na2O+2Fe For above reaction, if you are provided with 230g Na and 320g Fe2O3, then limiting reactant is	A. , Na B. Na2O C. Fe2O3 D. none of these
16	Which of the following is a limitation of balanced chemical equation	A. Conditions and rate of reactions B. Physical state and mechanism C. Reactants and products and their coefficients D. Both (a) and (b)
17	Moles of protons in 20g of SO3	A. 10 B. 20 C. 40 D. 80
18	Which of the following compound have empirical formula, but no molecular formula	A. H20 B. C6H6 C. H₂O₂ D. NaCl
19	If empirical formula of a compound is CH2 and its molecular mass is 56amu. What will beits molecular formula	A. CH2 B. C3H6 C. C2H4 D. C4H8
20	Atoms having same mass number but different atomic numbers are called.	A. Isotopes B. isobars C. Isotones D. isomers