

MDCAT Chemistry Chapter 11 S and P Block Elements Online Test

Sr	Questions	Answers Choice
1	No of unpaired electrons are maximum in	A. V+3 B. Mn+2 C. Fe+3 D. Cr+3
2	Which of the following pair has the same no. of electrons in d- subshell	A. Sc+3,Ti+3 B. Mn+2,Fe+3 C. Ti+3,V+3 D. Cr+3.Co+2
3	Which of these has at least one d electron	A. Sc+3 B. Mn+7 C. Ti+4 D. Cr+3
4	The total number of 3d-series transition elements is	A. 10 B. 40 C. 14 D. 58
5	Which ion has maximum number of unpaired electrons in 3d subshell and shows maximum paramagnetic behavior?	A. Cr+3 B. Ni+2 C. Co+2 D. Fe+3
6	Which of the elements has seven electrons in d-subshell?	A. Zn B. Co C. Cu D. Fe
7	d-d transition cannot be observed in	A. Cr B. Cu C. Mn D. Zn
8	which of the following is a typical transition metal?	A. Sc B. Y C. Ra D. Co
9	which one pair has the same oxidation state of-Fe?	A. FeSO ₄ and FeCl ₄ B. FeCl ₄ and FeCl ₃ C. FeSO ₄ and FeCl ₂ D. Fe ₂ (SO ₄) ₃ and FeSO ₄
10	Which pair of transition elements shows abnormal electronic configuration?	A. Sc and Zn B. Cu and Sc C. Zn and Cu D. Cu and Cr
11	Oxidation state of Mn' in KMnO ₄ . K ₂ MnO ₄ , MnO ₂ and MnSO ₄ is in the order	A. +7.+6.+2,+4 B. +6,+7,+2,+4 C. +7. +6.+4.+2 D. +4, +6, +7,+2
12	which of the following d blocks element can showthe highest oxidation number is its ompound	A. Chromium B. iron C. Copper D. Manganese
13	In [Ti (H ₂ O)] ₃ ⁺ which colour is transmitted	A. Yellow B. Blue and red C. Blue and yellow D. red and yellow
14	In the electronic configuration of Cr one electron from 4s sub-shell is transferred to 3d sub-shell because	A. The 3d orbital is of lower energy than 4s B. The half-filled d-subshell is more stable than 4 electrons having d-subshell C. The 4s orbital is of equal energy to 3d orbital D. 6 unpaired electron make Cr more paramagnetic

15	Which of the following is a non-typical transition element?	A. Cr B. Zn C. Mn D. Fe
16	When light is exposed to a typical transition element, then electrons jumps from low orbitals to higher orbitals in	A. f-orbitals B. s-orbitals C. p-orbitals D. d-orbitals
17	Highest oxidation state af the transition elements is	A. +8 B. +7 C. +5 D. +1
18	Electrons in 5d energy level are filled up in case of	A. Lanthanides B. Transition metals C. Actinides D. Rare gases
19	Zn has	A. Zero unpaired electrons B. Three unpaired electrons C. Five unpaired electrons D. One paired electrons
20	The number of unpaired electrons present in Fe ions is	A. 1 B. 2 C. 5 D. 0