

MDCAT Chemistry Chapter 11 S and P Block Elements Online Test

Sr	Questions	Answers Choice
1	[Ti (H ₂ O) ₆] ³⁺ ion is in colour.	A. Yellow B. Blue C. Violet D. Red
2	The energy difference of d-orbitals varies from	A. Atom to atom B. Ion to ion C. Electron to electron D. proton to proton
3	Which of the following are responsible for the colour developed in transition elements compounds?	A. s-orbitals B. p-orbitals C. d-orbitals D. f-orbitals
4	Transition compounds which occur as tripositive ions have no	A. 4s-electron B. 3p-electron C. 3s-electron D. 2s-electron
5	TiCl ₄ is used as catalyst for manufacture of	A. Sulphuric acid B. Plastics C. Ethanol D. Tetraethyl lead
6	Catalyst used for ammonia synthesis is	A. Cu B. Co C. Zn D. Fe
7	When light is exposed to transition element, then electrons jump from lower orbitals to higher orbitals in	A. Orbitals of s-subshell B. Orbitals of d-subshell C. Orbitals of p-subshell D. between different shells
8	d-d transition cannot be shown by	A. Cu ⁺¹ B. Sc ⁺³ C. Zn ⁺² D. All
9	Ti ⁺³ shows minimum absorption (maximum transmittance) at-----and-----wavelength	A. Yellow, Green B. Red, Yellow C. Blue, Green D. Red, Blue
10	Number of electrons involved in d-d transition of [Ti(H ₂ O) ₆] ³⁺	A. 1 B. 3 C. 2 D. 4
11	Which of the following transition metal forms colourless compounds in +4 oxidation state?	A. Ti B. Cr C. Cu D. Zn
12	Which of the following compound is expected to be colored	A. Na ₂ SO ₄ B. ZnCl ₂ C. MgF ₂ D. CuF ₂
13	The highest oxidation state of manganese is	A. +7 B. -7 C. +6 D. +4
14	The oxidation state of transition elements is usually	A. Variable B. Single C. Constant D. Infinite
15	Zinc does not show variable oxidation state, because	A. Its d-subshell is incomplete B. Its d-subshell is complete C. It is relatively soft metal D. It has two electrons in outermost

		shell
16	At which oxidation state Cu achieves electronic configuration of Zn ⁺²	A. 0 B. +2 C. +1 D. +3
17	The element which shows highest binding energy	A. V B. T C. So D. Cr
18	Which of the followings has electronic configuration of Ar in +3 oxidation state	A. Sc B. Mn C. Ti D. Zn
19	What is the sequence of electron take up and removal from 4s orbital a transition metal in 3d series?	A. Enters first, leaves after 3d electrons removal B. Enters after 3d electrons, leaves after 3d electrons C. Enters after 3d electrons, leaves first D. Enters first and leaves first
20	Variable Oxidation state of is related to transition elements	A. empty d-subshells B. Completely filled C. Partially filled d-subshell D. d-d transition