

Physics FSC Part 2 Chapter 20 Online MCQ's Test

Sr	Questions	Answers Choice
1	The line radiations emitted from by hydrogen filled discharge tube can be analyzed into.	A. Band spectrum B. Line spectrum C. Continuous spectrum D. Absorption spectrum
2	Balmer Empirical formula explains the electromagnetic radiation of any excited atom in terms of their.	A. Energy B. Mass C. Wave length D. Momentum
3	For Paschen series, the value of 'n' starts from	A. 2 B. 4 C. 6 D. 8
4	Which series lies in the ultraviolet region.	A. Balmer series B. Bracket series C. Ptund series D. Lyman series
5	Balmer series lies in	A. Visible region B. Invisible region C. Ultraviolet region D. Infrared region
6	The shortest wave length is Bracket series has wave length.	A. $16/R_n$ B. $R_n/16$ C. $16 R_n$ D. $4 R_n$
7	Balmer series lies in region of electromagnetic spectrum.	A. Infrared B. Visible C. Ultraviolet D. Fra infrared
8	Laser is a device which can produce:	A. Intense beam of light B. Coherent beam of light C. Monochromatic beam of light D. All of the above
9	Target material used in x-rays tube have following properties.	A. High atomic number and high melting point B. High atomic number and low melting point C. Low atomic number and low melting point D. High atomic number only
10	Frequency of x-rays depends upon.	A. Number of electrons striking target B. Accelerating potential C. Nature of the target D. Both B and C
11	When an electron in an atom goes from a lower to higher orbit its:	A. K.E increases , P.E decreases B. K.E increases , P.E increases C. K.E decreases , P.E increases D. K.E decreases , P.E decreases
12	In the Bohr's model of the hydrogen atom, the lowest orbit corresponds to:	A. Infinite energy B. Maximum energy C. Minimum energy D. Zero energy
13	In according with Bohr's theory the K.E of the electron is equal to:	A. $\frac{ke^2}{2r}$ B. $\frac{Ze^2}{r}$ C. $\frac{Ze^2}{r^2}$ D. $\frac{Ze^2}{2r^2}$
14	The Balmer series is obtained when all the transition of electrons terminate on	A. 1 st orbit B. 2 nd orbit C. 3 rd orbit D. 4 th orbit

15	For an atom of hydrogen atom the radius of the first orbit is given by:	B. $\frac{me}{4h^2}$ C. $\frac{h^2}{4\pi^2 kme}$ D. $\frac{h^2}{me}$
16	If 13.6 eV energy is required to ionize the hydrogen atom, then the required energy to remove an electron from $n=2$ is:	A. 10.2 eV B. 0 eV C. 3.4 eV D. 6.8 eV
17	1 rem =	A. 0.001 SV B. 0.01 SV C. 0.1 SV D. 1.01 SV