Chemistry Fsc Part 1 Chapter 9 Online Test

Sr	Questions	Answers Choice
1	Which one of the following salts do not hydrolyses	A. CuSO4 B. Na2CO3 C. NaCl D. AlCl3
2	Which one of the following salts dissolved in water to form a solution with a pH lesser than 7	A. NaCl B. CuSO4 C. Na2CO3 D. NH4Cl
3	Which one of the following salts dissolved in water to form a solution with a pH greater than 7	A. NaCl B. CuSO4 C. Na2CO3 D. NH4Cl
4	Which one of the following salts dissolved in water to form a solution with a pH greater than 7	A. NaCl B. CuSO4 C. Na2CO3 D. NH4Cl
5	Salt of weak acid with strong base when dissolved in water gives.	A. Acidic solution B. Basic solution C. Neutral solution D. None of above
6	When an ionic compound is dissolved in water, it dissociate into positive and negative ions, which are surrounded by H2O molecule, This process is known as.	A. Hydrolysis B. Hydration C. Saturation D. solvolysis
7	When an ionic compound is dissolved in water, it dissociate into positive and negative ions, which are surrounded byH2O molecule, This process is known as.	A. Hydrolysis B. Hydration C. Saturation D. solvolysis
8	Heat of solution of an ionic compound is equal to.	A. Hydration energy B. Lattice energy C. Slum of both 'a' and 'b' D. Difference of both a and b
9	Colligative properties are used to determine the	A. Freezing points B. Boiling point C. Atomic mass of an element D. Molar mass of solute
10	The vapour pressure of an aqueous solution of glucose is.	A. Equal to vapour pressure of water B. Independent of temperature C. More than vapour pressure of pure water D. Less than vapour pressure of pure water
11	Molal boiling constant for water is 0.52 ^{o}C. If 6 g of urea is dissolved in 100 g of water, what will be its boiling point.	A. 100.52 ^oC B. -100.52^oC C. 100^oC D. 99^oC
12	Molal boiling point elevation depends upon	A. Nature of solvent B. Natrue of solute C. Vapour pressure of solution D. None of these
13	The molal boiling point constant is the ratio of elevation of boiling point to	A. Molarity B. Mole fraction of solvent C. Molality D. Mole fraction of solute
14	Solubility of which substance decreases by increasing temperature.	A. NaNO3 B. KNO2 C. NaCl D. Ce2(SO4)3
15	Solubility of which substance decreases by incressing temperature.	A. NaNO3 B. KNO2 C. NaCl

		D. Ce2(SO4)3
16	Solubility curve of CaCl26H2O shows	A. Decrease in solubility with increase of temperature B. Increase in solubility with increase of temperature C. Discontinuous increase in solubility with temperature D. No effect of temperature on solubility
17	Water and Phenol are partially miscible to each other at room temperature when both liquids are mixed together which is upper layer.	A. Water in Phenol B. Phenol and water C. Pure phenol D. Pure water