Chemistry Fsc Part 1 Chapter 8 Online Test

Sr	Questions	Answers Choice
1	The pH of human blood is	A. 7.0 B. 7.4 C. 4.0 D. 6.5
2	K_a and K_b of a conjugate acid and are related with k_w as	A. $K_a + K_b = K_w$ B. $K_a - K_b = K_w$ C. $K_a \cdot K_b = K_w$ D. $K_a / K_b = K_w$
3	The effect of temperature on equilibrium was studied by	A. Lewis B. Van der wall C. Arrhenius D. Vant hoff
4	Le-Chatelier Braun principle is sometimes known as	A. Law of mass action B. Law of mobile equilibrium C. Law of active mass D. All of these above
5	The relationship between K_p and K_c is given by	
6	If the volumes of reactants and products are same in a gaseous phase reaction, then the equilibrium state is not affected by	A. Change of temperature B. Change of pressure C. Change of concentration D. Catalyst
7	Question Image	A. $dm^{+6} mole^{-2}$ B. $mole^2 dm^{-6}$ C. Mole dm^{-3} D. Having no units
8	At equilibrium stage of chemical reaction	A. The concentration of reaction is equal to concentration of products B. The rate constant of forward reaction is equal to rate constant of backward reaction C. The rate of forward reaction is equal rate of backward reaction D. The energy of activation of forward step is equal to energy of activation of backward step
9	An excess of aqueous silver nitrate is added to aqueous barium chloride and precipitate is removed by filtration. What are the main ions in the filtrate	
10	The solubility product of AgCl is $2.0 \times 10^{-10} mol^2 dm^{-6}$ The maximum concentration of Ag^+ ions in the solution is	A. 2.0×10^{-10} mol dm^{-3} B. 1.41×10^{-5} mol dm^{-3} C. 1.0×10^{-10} mol dm^{-3} D. 4.0×10^{-20} mol dm^{-3}
11	The pH of 10^{-3} mol dm^{-3} of an aqueous solution of H_2SO_4 is	A. 3.0 B. 2.7 C. 2.0 D. 1.5
12	Question Image	A. The value of K_p falls with a rise in temperature B. The value of K_p falls with increasing pressure C. Adding V_2O_5 catalyst increase the equilibrium yield of sulphur trioxide

D. The value of K_p is equal to K_c

13   For which system does the equilibrium constant K_c has units of (concentration)$^{-1}$