Chemistry Fsc Part 1 Chapter 8 Online Test

Sr	Questions	Answers Choice
1	The sum of pH and pOH is	A. 0 B. 7 C. 14 D. 10
2	When concentration of one product is removed at equilibrium stage, in which direction it moves to reestablish equilibrium.	A. Forward B. Reverse C. Neither forward nor reverse D. Equally move in both direction
3	When solid KI dissolved in water, its heat of solution is positive. What would happen to dissolution when temperature is increased.	A. Increases B. Decreases C. Remain same D. Firs increases than decreases
4	For the equilibrium system $N_2 + O_2 + Heat = 2NO$ the equilibrium constant deceases by	A. Decreasing the temperature B. Adding a catalyst C. Adding N_2 D. Adding NO
5	the substance which increase the rate of reaction but remains unchanged at the end of the reaction is called.	A. Indicator B. Promoter C. Catalyst D. Activated complex
6	Almost forward reaction is complete when value of Kc is	A. very high B. Very small C. Neither large nor very small D. No correlation
7	The unit of Kc for the reaction $N_2 + O_2 = 2NO$ will be	A. mol dm-3 B. mol-1 dm+3 C. mol-2 dm+6 D. No units
8	A chemical reaction A ---------- B is said to be in equilibrium when	A. Complete conversion of A to B has taken place B. Conversion of A to B is 50% complete C. Rate of transformation of A to B is equal to B to A D. 50% Reactant have been changed to B
9	The pH of 10-3 mole dm-3 of an aqueous solution of H2SO4 is.	A. 3.0 B. 2.7 C. 2.0 D. 1.5
10	Catalyst used in preparation of NH3 from N2 and H2 is.	A. Ni B. Fe C. Pt D. V2O5
11	Optimum pressure in Haber's process for synthesis of Ammonia is	A. 100 -150 atm B. 200- 300 atm C. 350 - 450 atm D. 500 - 600 atm
12	_____ was derived by C.M Guldberg and P Waage in 1864	A. Law of conservation of Mass B. Law of mass action C. Law of conservation of energy D. Distribution law
13	The law of mass action was given by	A. D.C. down and P wage B. Gay Lussic and C.M C. C.M Goldberg and P. Waage D. Hendeson and Le Chateller's
14	The pK_a value of CH_3COOH is 4.74 when we mix CH_3COOH and CH_3COONa in the ratio of 10:1, tehn the pH of the buffer is	A. 4.74 B. 5.74 C. 3.74 D. 7.00

15	The number of moles of acid or base required by one dm^3 of buffer to alter its pH by one unit is called	A. Buffer efficiency B. Buffer capacity C. Buffer action D. None
16	A buffer solution can be prepared by mixing	A. Weak acid and its salt with weak base B. Weak base and its salt with strong acid C. Strong acid and its salts with weak base D. Strong base and its salt with weak acid
17	The value of pH and P^{OH} of pure water at 25° C is	A. 14 B. 7 C. 1×10^{-14} D. 1×10^{14}