Chemistry Fsc Part 1 Chapter 7 Online Test

Sr	Questions	Answers Choice
1	The born Haber cycle is the best application of law.	A. Boyle's B. Dalton's C. Hess's D. Graham's
2	The pressure of oxygen inside the bomb calorimeter is.	A. 100 atm B. 50 atm C. 25 atm D. 20 atm
3	The change in heat energy of a chemical reaction at constant temperature and pressure is called.	A. Enthalpy change B. Bond energy C. Heat of sublimation D. Internal energy change
4	The net heat change in a chemical reaction is same whether it is brought about in two or more different ways in one or several steps. It is known as	A. Henry's law B. Hess's law C. Joule's principle D. Law of conservation of energy
5	For a given process, the heat chagnes of constant pressure and at constant volume are related to each other as.	A. qp= qv B. qp<qv C. qp>qv D. None of these
6	In endothermic reactions, the heat content of the.	A. Products is more than that of reactants. B. Reactants is more than that of products C. Both a and b D. Reactants and products are equal
7	If an endothermic reaction is allowed to take place very rapidly i the air, the temperature fo the surrounding air.	A. Remains constant B. Increase C. Decrease D. Remain unchanged
8	Whenever a reaction is endothermic, then it means that	A. Heat is transferred from surrounding to the system B. Heat is transferred system to the surrounding C. Heat content of the product is greater than that of reactants D. Heat content of the reactants is greater than the products
9	One of the following statements about Born-Haber cycle is correct. Which is that statement	A. Born-Haber cycle is different from Hess's law B. The energy change in a cyclic process is not zero C. The lattice energy of the crystalline substances can be calculated easily D. Heat of formation of the product and the lattice energy of the substance can be calculated simultaneously
10	The net change in a chemical reaction is same whether it takes place directly or indirectly is	A. Henry's law B. Charles's C. Hess's lass D. Graham's law
11	The S.I. units for the molar heat capacity are	A. Joule Cm⁻³degree⁻¹ B. Joule deg⁻¹atm⁻¹ C. Joule deg⁻¹mol⁻¹ D. Joule deg⁻¹kg⁻¹
12	Question Image	A. Enthaply change of the system

13	At constant volume of a system remains constant and the heat is absorbed by the system, them amount of heat absorbed is called	B. Internal energy change of the system C. Total enthalphy of the system D. Total internal energy of the system
14	What happens to the enthalpy change when the coeffficients of a chemical equation are douubled	A. It doubles B. It becomes half C. It does not change D. It cannot be predicted
15	Most of the reaction which give stable products are	A. Exothermic B. Endothermic C. Isothermal D. None of these
16	A material or a collection of materials which is under study is called	A. State function B. Degree and joule C. Degree and ergs D. Calorie and joule
17	The net heat change in a chemical reaction is same whether it is brought about in two or more different ways in one or several steps. It is knnown as	A. Henry's law B. Hess's law C. Joule's principle D. Law of combustion