

Chemistry Fsc Part 1 Chapter 6 Online Test

Sr	Questions	Answers Choice
1	Which element has highest ionization potential.	A. Li B. B C. Be D. C
2	Which of the following molecule obey octet rule.	A. BF ₃ B. BCl ₃ C. NH ₂ D. SF ₆
3	Which compound does not obey the octet rule.	A. NH ₃ B. BCl ₃ C. H ₂ O D. CH ₄
4	SP ³ hybridization is important in describing the bonding in	
5	If we want to change O ₂ to O ₂ ⁻¹ The electron is to be placed in	
6	VSEPR theory was proposed by	A. Nyholm and Gillespie B. Kossel C. Lewis D. Sidwick
7	The shape of H ₃ O is	A. Tetrahedral B. Angular C. Trigonal planer D. Pyramidaj
8	Which of the following molecules has a co-ordinate covalent bond	A. NH ₄ ⁺ Cl ⁻ B. NaCl C. HCl D. AlCl ₃
9	The nature of bond in diamond is	A. Electrovalent B. Metallic C. Coordinate covalent D. Covalent
10	Fluorine molecule (F ₂) is formed by the overlap of	A. s - s orbital B. s - p orbital C. p - p head on overlapping of orbitals D. p - p parallel overlapping of orbitals
11	Which of the following is true for ionic compounds	A. They are non-electrolytes in the molten state B. They have bonds which are directional C. They conduct electricity in solid state D. They are generally more soluble in polar solvents than in non-polar solvents
12	Those elements whose electronegativities are 1.2 and 3.2, react to form	A. Ionic bond B. Covalent bond C. Gaseous substance D. Definiting a liquid substance
13	The shielding effect of the inner electrons is responsible for	A. Increasing ionization energy values B. Decreasing ionization energy values C. Increasing electron affinity D. Increasing electonegativity
14	The octet rule is not followed in the formation of	A. NF ₃ B. CF ₄ C. CCl ₄ D. PCl ₅
		A. O ₂ ²⁺

15	Which of the following species has unpaired electrons in anti-bonding molecular orbitals	B. N_2^{+} C. B_2 D. F_2
16	Which of the hydrogen halides has the highest percentage of ionic character	A. HF B. HBr C. HCl D. HI
17	Which of the following molecules has zero dipole moment	A. NH_3 B. CHCl_3 C. H_2O D. BF_3