Chemistry Fsc Part 1 Chapter 5 Online Test

Sr	Questions	Answers Choice
1	Lyman series lies in spectral region	A. Infrared B. Ultra violet C. Visible D. None of these
2	The velocity of photon is.	A. Depends on its source B. Equal to square of its amplitude C. Depends on its wavelength D. Independent of its wavelength
3	In the ground state of an atom, the electrons is present.	A. In the nucleus B. In the second shell C. Nearest to the nucleus D. Farthest form the nucleus
4	Bohr's model of atom is contradicted b y.	A. Planck quantum theory B. Quantization of energy of electrons C. Heisenberg's uncertainty principle D. Quantization of angular momentum
5	Rutherford's model of atom failed because.	A. The atom did not have a nucleus and electron B. It did not account for the attraction between protons and neutrons. C. It did not account for stability of the atom D. There is actually no space between the nucleus and the electrons.
6	When fast neutron carries nuclear reaction with nitrogen it ejects aprticles.	A. Alpha B. Gamma C. Beta D. Nil
7	The e/m value for the positive rays in maximum for the gas.	A. Helium B. Oxygen C. Nitrogen D. Hydrogen
8	Positive rays were discovered by.	A. J.J.Thomson B. Goldstein C. Ruther ford D. William Crookes
9	The nature of positive rays depends on	A. The nature of electrode. B. The nature of discharge tube C. The nature of residual gas D. All of the above
10	Cathode rays strike alumina and produce a .....colour.	A. Red B. Blue C. Yellow D. Green
11	Bohr's model of atom is contradicted by	A. Planck quantum theory B. Quartization of energy of electrons C. Heisenberg's uncertainty principle D. Quartization of angular members
12	The charge on electron was determined by millikan in his oil drop experiment and its value is	A. 6.023 x 10⁻²³C B. 1.602 x 10⁻²³C C. 1.602 x 10⁻¹⁹C D. 6.625 x 10⁻³⁴C
13	Which of the following represents electronic configuration of the most electropositive elements	A. He [2s⁻¹] B. Xe [6s¹] C. He [2s²] D. Xe [6s²]
14	Which of the following particles would on losing an electron has its outermost p-orbital as half filled	A. Nitrogen atom B. O⁺ion C. P⁻¹ion D. S⁺¹ion

15	The azimuthal quantum number / = 2, then M(Magnetic quantum number) can have values as	A. +1, -1 B. +1, 0, -1 C. +2, +1, 0,1, -1, 2 D. +3, +2, +1, 0, 1, -2, -3
16	If uncertainty position of an electron is zero, the uncertainty in its momentum would be	A. Zero B. Infinite C. Both a and b D. None of these
17	The line of the balmer series in the visible region of the spectrum, but the limiting line, in the series lies in	A. Visible region B. X-Ray region C. I.R region D. U.V. region