

## Chemistry Fsc Part 1 Chapter 5 Online Test

Sr	Questions	Answers Choice
1	Quantum number values for 2p orbitals are	A. $n = 2, l = 1$ B. $n = 1, l = 2$ C. $n = 1, l = 0$ D. $n = 2, l = 0$
2	In the ground state of an atom the electron is present.	A. In the nucleus B. In the second shell C. Nearest to the nucleus D. Farthest from the nucleus
3	Splitting of spectral lines when atoms are subjected to strong electric field is called.	A. Zeeman effect B. Stark effect C. Photoelectric effect D. Compton effect
4	Bohr's model of atom, is contradicted by.	A. Planck quantum theory B. Pauli's exclusion principle C. Heisenberg's uncertainty principle D. All of the above
5	Rutherford's model of atom failed because	A. The atom did not have a nucleus and electrons. B. It did not account for the attraction between protons and neutrons. C. It did not account for the stability of the atom D. There is actually no space between the nucleus and the electrons.
6	The wave number of the light emitted by a certain source is $2 \times 10^6 \text{ m}^{-1}$ . The wavelength of this light will be.	A. 500 nm B. 5000 nm C. 200 nm D. $5 \times 10^{-7} \text{ m}$
7	the nature of the positive rays depend on	A. The nature of the electrode B. The nature of the discharge tube C. The nature of the residual gas D. All of the above
8	When 5d orbital is completed then entering electron goes into.	A. 6s B. 6p C. 6d D. 6f
9	$n+1$ value of 6d orbital is.	A. 08 B. 09 C. 10 D. 18
10	When 6d orbital is complete, the entering electron goes into.	A. 7f B. 7s C. 7p D. 7d
11	An orbital which is spherical and symmetrical is	A. s-Orbital B. p - Orbital C. d- Orbital D. f - Orbital
12	Orbitals having same energy are called.	A. Hybrid orbitals B. Valance orbitals C. Degenerate orbitals D. d- orbitals
13	Maximum number of electrons in f-subshell is.	A. 2 B. 6 C. 10 D. 14
14	The electron in a subshell is filled according to formula.	A. $2n^2$ B. $2(2l+1)$ C. $(2l+1)$ D. None of these

15	Quantum number values for 2p orbitals are.	A. $n = 2, l = 1$ B. $n = 1, l = 2$ C. $n = 1, l = 0$ D. $n = 2, l = 0$
16	When atoms are subjected to strong electric field, splitting of spectral lines is called.	A. Zeeman effect B. Stark effect C. Photoelectric effect D. Compton effect
17	The wave number of the light emitted by a certain source is $2 \times 10^6 \text{ m}^{-1}$ . The wave length of this light is.	A. 500 nm B. 500 m C. 200 nm D. 600 m