

Chemistry Fsc Part 1 Chapter 4 Online Test

Sr	Questions	Answers Choice
1	The polarizabilities of elements mostly increase down the group due to the reason that	A. The atomic numbers increase B. Number of protons increase C. Number of shells increase along, with increase of shielding effect D. The behavior of the elements remain the same
2	The repulsion of electronic clouds of the molecules are responsible for the attractive forces among the molecules. These forces are	A. Dipole-induced dipole forces B. Ion-dipole forces C. Instantaneous dipole-induced dipole forces D. Dipole-dipole forces
3	In order to mention the B.P. of water at 110°C, the external pressure should be	A. Between 760 torr and 1200 torr B. Between 200 torr and 760 torr C. 760 torr D. Any value of pressure
4	When water freezes at 0°C, its density decreases due to	A. Cubic structure of ice B. Empty spaces present in the structure of ice C. Change of bond lengths D. Change of bond angles
5	NH ₃ shows a maximum boiling point among the hydrides of V-A group elements due to	A. Very small size of nitrogen B. Lone pair electrons present on Nitrogen C. Enhanced electronegative character of Nitrogen D. Pyramidal structure of NH ₃
6	Acetone and chloroform are soluble in each other due to	A. Intermolecular hydrogen bonding B. Dipole-dipole interaction C. Instantaneous dipoles D. All of the above
7	London dispersion forces are the only forces present among the	A. Molecules of water in liquid state B. Atoms of helium in gaseous state at high temperature C. Molecules of solid iodine D. Molecules of hydrogen chloride gas