

Chemistry Fsc Part 1 Chapter 4 Online Test

Sr	Questions	Answers Choice
1	In cubic and hexagonal closest packing which layer has different arrangement.	A. First B. Second C. Third D. Fourth
2	When external pressure is 23.7 torr, boiling point of water is	A. 100 °C B. 200 °C C. 98 °C D. 25 °C
3	Vapour pressure of a substance does not depend upon.	A. Temperature B. Intermolecular forces C. Surface area D. Physical state of water
4	Ice float over water because.	A. Its structure is diamond like B. Its density is maximum at 4 °C C. It is less dense than water D. It has no regular arrangement of molecules.
5	When liquid water changes to ice its volume expands. The expansion in volume is.	A. 5% B. 9% C. 10% D. 18%
6	Conductivity of metal decreases by increasing temperature because.	A. Atoms are converted to ions B. Atoms oscillates and hinder the movement of free electrons. C. Ions are converted into atoms D. Velocity of mobile electrons increases
7	Which pair of molecule have Debye forces in them	A. Ar and Ar B. Argon and water C. Na ⁺ ions and water D. water and water
8	Which one of the following inter molecular forces are present in neon gas molecules.	A. Hydrogen bond B. dipole -Dipole attraction C. London dispersion force D. Hydrogen bonding and London dispersion force
9	Which of the following elements in its crystalline form will have the lowest enthalpy change of vaporization	A. Chlorine B. Argon C. Phosphorus D. Silicon
10	Long chains of amino acids are coiled about one another into a spiral by	A. Covalent bond B. Ionic bond C. Hydrogen bond D. Van Der Waal's forces
11	Exceptionally low acidic strength of HF is due to.	A. Strong polar bond between H and F B. Smaller size of fluorine C. Strong hydrogen bonding D. electronegativity of fluorine
12	Down the VII -A group, polarizability generally.	A. Increases B. Decreases C. Remain constant D. Negligible
13	Diamond is a bad conductor because.	A. It has tight structure. B. It has a high density C. There is no free electron present in the crystal of diamond to conduct electricity D. None of the above
		A. CaF ₂ B. Si

14	Which of the following is psuedo solid	B. Glass C. NaCl D. All
15	The molecule of CO ₂ in dry ice form the.	A. Ionic crystals B. Covalent crystals C. Molecular crystals D. Any type of crystals
16	Amorphous solids.	A. Have sharp melting points. B. Undergo clean cleavage when cut with knife C. Have perfect arrangement of atoms D. Can possesses small regions of orderly arrangement of atoms.
17	Ionic solids are characterized by.	A. Low melting points B. High vapour pressures C. Good conductivity in solid state D. Solubility in polar solvents