

Chemistry Fsc Part 1 Chapter 3 Online Test

Sr	Questions	Answers Choice
1	Gases deviate from ideal behaviour at high pressure. Which of the following is correct for non-ideal behaviour of gases	<p>A. At high pressure, the gas molecules move in one direction only</p> <p>B. At high pressure, the collisions between the gas molecules are increased</p> <p>C. At high pressure, the volume of the gas becomes insignificant</p> <p>D. At high pressure, the intermolecular attraction becomes significant</p>
2	Equal masses of methane and oxygen are mixed in an empty container at 25°C, the fraction of total pressure exerted by oxygen is	<p>A. 1/3</p> <p>B. 8/9</p> <p>C. 1/9</p> <p>D. 16/17</p>
3	the order of the rate of diffusion of gases NH ₃ , SO ₂ , Cl ₂ and CO ₂ is	<p>A. NH₃ < SO₂ < Cl₂ < CO₂</p> <p>B. NH₃ < CO₂ < SO₂ < Cl₂</p> <p>C. Cl₂ < SO₂ < CO₂ < NH₃</p> <p>D. NH₃ < CO₂ < Cl₂ < SO₂</p>
4	The molar volume of CO ₂ is maximum at	<p>A. STP (0°C and 1 atm)</p> <p>B. 127°C and 1 atm</p> <p>C. 0°C and 2 atm</p> <p>D. 273°C and 2 atm</p>
5	How should the condition be changed to prevent the volume of a given gas from expanding when its mass is increased	<p>A. Temperature is lowered and pressure is increased</p> <p>B. Temperature is increase and pressure is lowered</p> <p>C. Temperature and pressure both are lowered</p> <p>D. Temperature and pressure both are increased</p>
6	If absolute temperature of a gas is doubled and the pressure is reduced to one half, the volume of the gas will	<p>A. Remain unchanged</p> <p>B. Increase four times</p> <p>C. Reduce to 1/4</p> <p>D. Be doubled</p>
7	Which of the following will have the same number of molecules at STP	<p>A. 280 cm³ of CO₂ and 280 cm³ of N₂O</p> <p>B. 11.2 dm³ of O₂ and 32 g of O₂</p> <p>C. 44 g of CO₂ and 11.2 dm³ of CO</p> <p>D. 28 g of N₂ and 5.6 dm³ of oxygen</p>
8	The number of molecules in one dm ³ of water is close to	
9	Pressure remaining constant at which temperature the volume of a gas will become twice of what it is at 0°C	<p>A. 546°C</p> <p>B. 200°C</p> <p>C. 546 K</p> <p>D. 273 K</p>