Chemistry Fsc Part 1 Chapter 3 Online Test

Sr	Questions	Answers Choice
1	The partial pressure of oxygen in lungs is	A. 760 torr B. 320 torr C. 159 torr D. 116 torr
2	The commonly used unit of pressure by meteorologists is.	A. Atmosphere B. Pascal C. Milli D. Pound inch3
3	Vapour pressure of liquid depends upon	A. Amount of liquid B. Surface area C. Temperature D. Size of container
4	Partial pressure of oxygen in the air is.	A. 156 torr B. 157 torr C. 158 torr D. 159 torr
5	Pressure remaining constant at which temperature the volume of a gas will come twice of what it is at 0 ^{o}C	A. 546 ^oC B. 200 ^oC C. 546 K D. 273 K
6	Mass of 22.4 dm3 of N2 at STP is.	A. 28 gm B. 14 gm C. 1.4 gm D. 2,8 gm
7	The molar volume of CO2 is maximum at.	A. STP B. 127 ^oC and 1 atm C. 0 ^oC and 2 atm D. 273 ^oC and 2 atm
8	If absolute temperature of the gas is doubled and the pressure is reduced to one half the volume the gas will.	A. Remains uncharged B. Increase four time C. Reduce to 1/4 D. Be doubled
9	Borax has the chemical formula.	A. KNO3 B. Na2B3O7.10H2O C. Na2CO3 D. NaNO3
10	Noon has low critical temperature and pressure as compared to other gases. the most probable reason is that	A. Its octet is complete B. It is a monoatomic gas C. It has very low polarizability D. It has least forces of attraction
11	The highest temperature above which a gas cannot be liquified, no matter how much the pressure is applied is known as	A. Boiling temperature B. Consulate temperature C. Absolute zero D. Critical temperature
12	The molecules of a gas show more deviation from ideal behaviour at low temperature, because	A. Attractive force dominate at low temperature B. Kinetic energies are increased C. Collisions become less frequent D. Densities of the gases increase
13	The free expansion of the gas from high pressure towards the low pressure causes	A. Increase of temperature B. Decrease of temperature C. Grater number of collisions among the molecules D. Decrease of velocity of gas molecules
14	Rate of diffusion of CO and N_2 are same at room temperature due to the reason, that	A. Both are diatomic molecules B. Both have same multiple bond in them C. Both have lone pairs in them D. Both have same molar masses

15	In gas occupies a volume of 2 dm^3 at 27°C and I atm pressure. The expression for its volume at S.T.P. is	
16	Gas equation is derived by combining	A. Avogadro's and Charles's Law B. Boyle's and Charles's Law C. Avogadro's and Boyle's Law D. Avogadro's, Boyle's and Charles's Law
17	Normal temperature and pressure (S.T.P) of gas rafers to	A. 273 K and 76 mm Hg B. 273° C and 760 mm Hg C. 273 K and 760 mm Hg D. 273° C and 76 mm Hg