Chemistry Fsc Part 1 Chapter 3 Online Test

Sr	Questions	Answers Choice
1	Observed pressure is less than ideal pressure for any gas due to	A. Intermolecular forces B. Size of molecules C. Boiling point of molecules D. Both a and c
2	What are the SI units of Van der Waal constant 'a'	A. atm dm3 mol-2 B. atm dm6 mol-2 C. Nm4 mol-2 D. Nm mol-1
3	What are the SI of excluded volume 'b' in Vander waal equation.	A. dm³ mol -1 B. m3 mol -1 C. mol dm-3 D. mol m-3
4	The Van der Waals' equation explain the behaviour of.	A. Ideal gas B. Real gas C. Vapours D. Non ideal gases
5	Critical temperature of CO2 gas is.	A. 31.1 ^oC B. 13.1 K C. 13.1 ^oC D. 1.31 ^oC
6	More ideal gas at room temperature is.	A. CO2 B. NH3 C. SO2 D. N2
7	A real gas can be liquefied if.	A. Temperature is more than critical temperature. B. Temperature is less than critical temperature C. Pressure is more than critical pressure and temperature is less than critical temperature D. Its pressure is less than critical pressure
8	An Ideal gas can not be liquefied because.	A. Its critical temperature is always above 0 ^oC B. It molecules are relatively smaller in size C. Its solidify before becoming a liquid D. Force operative between its molecules are negligible
9	Which one of the following gases cannot be liquefied by Line's method.	A. Water vapours B. NH3 C. Nitrogen D. H2
10	Gas molecules show more deviation from ideal behaviour at high pressure because.	A. Velocity of molecules increases B. Velocity of molecules decreases C. Force of attraction between molecules increases D. Force of collision per unit area increases
11	Under which conditions of temperature and pressure will a real gas behave most like an ideal gas.	A. Low temperature and how pressure B. High temperature and high pressure C. Low temperature and high pressure D. High temperature and low pressure
12	The concept of distribution of velocities among the gas molecules was given by.	A. Clausius B. Mexwell C. Boltzmann D. Vander waal

13	Which is not example of natural plasma.	A. Lightening bolt B. Aurora C. Neon sign D. Sum
14	The total kinetic energy of one mole of an ideal gas is given by	A. 3/2 RT B. 1/2 KT C. 1/2 RT D. 3/2 KT
15	Total pressure of mixture of two gases is.	A. The sum of their partial pressures. B. The difference of their partial pressures C. The product of their practical pressures D. The ratio of their partial pressures
16	Which mixture of gases is used by the deep sea divers.	A. Oxygen and nitrogen B. Oxygen and helium C. Oxygen and carbon di oxide D. Oxygen and water vapours
17	Gases exert pressure on the walls of the container becasue the gas molecules.	A. Collide with each other B. Collide with walls of container C. Have definite volume D. Obey the gas laws