Chemistry Fsc Part 1 Chapter 11 Online Test

Sr	Questions	Answers Choice
1	After 3 half lives of a chemical reaction, the % fraction of the amount left is	A. 6.25 B. 75 C. 12.5 D. 50
2	The unit of rate constant depends upon	A. Number of reactants B. Concentration terms C. Order of reaction D. Molecularity of reaction
3	A substance which rerareds the rate of a reaction is called	A. Inhibitor B. Activator C. Auto-catalyst D. None of these
4	The influence of temperature on reaction rate is predicated by	A. Free energy change of reaction B. Arrhenius equation C. Van der weal's equation D. Kinetic equation
5	If the energy of the activated complex lies close to energy of reactants, it means that reaction is	A. Slow B. Fast C. Exothermic D. Endothermic
6	The power of which the concentration of a substance appears in the rate expression is known as	A. order of reaction with respect to that substance B. Rate of reaction C. Order of reaction D. Molecularity of reaction
7	The true representation for the units of rate constant K for the first order reaction	A. sec^{-1} B. $mole\ dm^{-3}s$ C. $mole\ dm^{-3}s^{-1}$ D. $mole^{-1}dm^{+3}s^{-1}$
8	The unit of rate constant is same as that of rate of reaction in	A. first order reaction B. Second order reaction C. Third order reaction D. Zero order reaction
9	The mathematical relation between the rate of reaction and the concentrations of the reactants is known as the	A. Rate equation B. Rate law C. Arrhenius equation D. Both a and b
10	Velocity constant is the rate of reaction when the concentrations of reactants are	A. Zero B. Unity C. Two D. Three
11	The reaction that involves gases, its rate does not depend upon	A. Catalyst B. Temperature C. Moles dm^{-3} D. Partial pressure
12	When a reaction proceeds in a sequence of steps, the overall rate is determined by	A. Fastest step B. Slowest step C. Order of different steps D. Molecularity of all steps
13	The unit of the rate constant is same as that of the rate of reaction in	A. First order reaction B. Second order reaction C. Zero order reaction D. Third order reaction
14	With increase in 10°C temperature, the rate of reaction double. This increase in rate of reaction is due to	A. Decrease in activation energy of reaction B. Decrease in the number of collisions between reactant molecules C. Increases in activation energy of

		reactants
		D. Increase in number of effect collisions
15	The rate of reaction	A. Increases as the reaction proceeds B. Decreases as the reaction proceeds C. Remains the same as the reaction proceeds D. May decrease or increase as the reaction proceeds
16	Question Image	A. 1 B. 2 C. 3 D. None of these
17	In zero order reaction, the rate is independent of	A. Temperature of reaction B. Concentration of reactants C. Concentration of products D. None of these