

## Chemistry Fsc Part 1 Chapter 11 Online Test

Sr	Questions	Answers Choice
1	When a chemical reaction is completed.	A. Instantaneous rate > average rate B. Instantaneous rate = average rate C. Instantaneous rate is zero D. Both average and instantaneous rates become zero
2	Which one of the following statements is incorrect.	A. Enzymes are protein in nature B. Enzymes are catalyst C. Enzymes can catalyze any reaction D. Urease is an enzyme
3	The substance which decrease the activity of a catalyst is called.	A. Promoter B. Activator C. Inhibitor D. Positive catalyst
4	A substance which itself is not a catalyst but increases the activity of a catalyst is called.	A. Promoter B. Poisoner C. Inhibitor D. Enzyme
5	A type of metals are usually used as catalyst.	A. Coinage metal B. Alkali metals C. Transition metals D. alkaline earth metals
6	Which statement is incorrect about activated complex	A. Short lived B. Maximum energy C. Unstable combination of atoms D. Less energy than Ea
7	When a reaction occurs in many steps then the slowest step is.	A. Mechanism step B. Rate determining step C. enthalpy determining step D. None of the above
8	Which statement is incorrect about order of reaction	A. It cannot be determined experimentally B. It is determined experimentally C. Sum of exponents in rate equation D. It can have fraction value
9	Photosynthesis a photochemical reaction has order of reaction	A. 0 B. 1 C. 2 D. Fractional order
10	When a chemical reaction is completed then	A. Instantaneous rate > average rate B. Instantaneous rate = average rate C. Instantaneous rate is zero D. Both average and instantaneous rates become zero
11	a zero order reaction is one in which	A. Reactants do not react B. One reactant is in large excess C. Concentration of reactant does not change with passage of time D. Rate is affected by changing concentration of reactants
12	After 3 half life of a chemical reaction, amount of reactant unreactive will be.	A. 50% B. 25% C. 12.5% D. 6.25%
13	Half -Life for a given reaction is doubled if initial concentration is doubled. The order of reaction is.	A. 0 B. 1 C. 2 D. 3
14		A. Initial concentration of the compound

14	Half -Life period of a first order reacting is independent of.	B. Conditions of temperature C. Presence of catalyst D. All the above
15	Chemical reactivity of different substance is controlled by	A. Atomic number B. Electronic arrangement C. Mass number D. Number of isotope of reactant elements
16	If reactants have very low activation energy it means that reaction is.	A. Slow B. Fast C. Endothermic D. Exothermic
17	Rate of a chemical reaction generally increase rapidly even for small increase in temperature because of rapid increase in the	A. Collisions frequency B. Activation energy C. Average KE of molecules D. Fraction of molecules with energies more than activation energy