

## Chemistry Fsc Part 1 Chapter 10 Online Test

Sr	Questions	Answers Choice
1	In silver oxide battery, anode is made of.	A. Zinc B. Copper C. Lead D. Graphite
2	Which is not chargeable cell	A. Lead accumulator B. NiCAD cell C. Fuel cell D. Alkaline battery
3	In NICAD dry cell, the cathode and anode is made up of.	A. Ca and Ag B. Ni and CdO <sub>2</sub> C. NiO <sub>2</sub> and Cd D. Ag and Ag <sub>2</sub> O
4	Fuel cells are mostly used in space air crafts as the source of.	A. Power only B. Drinking water C. Drinking water and power D. Fuel and drinking water
5	Electrode of the lead storage battery are immersed in dilute H <sub>2</sub> SO <sub>4</sub> which has strength by mass	A. 100% B. 98% C. 30% D. 10%
6	In lead accumulator cathode is made up of.	A. Pb B. Pb coated with PbO <sub>2</sub> C. PbSO <sub>4</sub> D. Mixture of Pb and PbO <sub>2</sub>
7	Which is not use of electrochemical series.	A. Feasibility of reaction B. Measurement of EMF of cell C. Comparison of reactivity with water or acids D. Determination of atomic and ionic radii
8	Which has greater reduction potential	A. Na B. H <sub>2</sub> C. Zn D. F <sub>2</sub>
9	The over all positive value for the reaction potential predicts that process is energetically.	A. Not feasible B. Feasible C. Impossible D. No indication
10	Cell potential depends upon	A. Temperature B. Concentration of ions C. Nature of electrolyte D. All of above
11	The difference of potential of two electrodes when concentration of solution is 1 M each at 25 °C and 1 atmosphere is called.	A. Electrode potential B. Standard cell potential C. Cell reaction D. Cell voltage
12	Standard hydrogen electrode has an arbitrarily fixed potential	A. 0.00 volts B. 1.00 volt C. 0.10 volt D. None of these
13	In Daniel cell, if salt bridge is removed between the two half cells, the voltage.	A. Drops to zero B. Does not changes C. Increases gradually D. Increases rapidly
14	A cell in which electric current is produced as a result of spontaneous redox reaction is called.	A. Electrolytic cell B. Galvanic cell C. Half cell reaction D. Down's cell
15	Alkali and alkaline earth metal are usually obtained by	A. Decomposition of their carbonates B. By heating their hydroxide C. electrolysis of molten metal oxides D. electrolysis of molten metal chlorides

D. Electrolysis of molten metal halides

16 In electrolysis of aqueous NaCl, Cl-ions are.

- A. Oxidized at anode
- B. Oxidized at cathode
- C. Reduced at cathode
- D. Neither oxidized nor reduced

17 Electrolysis is used for

- A. Electroplating
- B. Refining of copper
- C. Manufacture of caustic soda
- D. All of the above