

## Chemistry Fsc Part 1 Chapter 1 Online Test

Sr	Questions	Answers Choice
1	A limiting reactant is one.	A. Which is present in least amount B. Which produces minimum number of moles of product C. Which produces maximum number of moles product D. Does not effect the amount of product.
2	If four moles of sulphur dioxide are oxidized to sulphur trioxide, how many moles of oxygen are needed.	A. 0.5 B. 1.0 C. 1.5 D. 2.0
3	When one mole of each of the following is completely burned in oxygen, which gives the largest mass of carbon dioxide.	A. Diamond B. C <sub>2</sub> H <sub>6</sub> C. Methane D. CO <sub>2</sub>
4	What is the mass of aluminium is 204 g of the aluminum oxide Al <sub>2</sub> O <sub>3</sub> .	A. 26 g B. 27 g C. 54 g D. 108 g
5	How many molecules of CO <sub>2</sub> are formed by burning 12 g carbon with excess of oxygen.	A. $3.01 \times 10^{23}$ B. $1 \times 10^{23}$ C. $6.02 \times 10^{23}$ D. $1.03 \times 10^{23}$
6	How many moles of water results by burning 4 mole of H <sub>2</sub> with excess of oxygen.	A. 1 mol B. 2 mol C. 3 mol D. 4 mol
7	Stoichiometric calculations cannot applied to reversible reactions because.	A. Product again changes to reactant B. Less product is formed C. Reaction goes only in one direction D. Products do not disappear.
8	The number of electrons in one mole of hydrogen gas is.	A. $6.02 \times 10^{23}$ B. $12.04 \times 10^{23}$ C. Only two D. Indefinite
9	The Avogadro constant is the number of.	A. Atoms in 1 g of helium ga B. Molecules in 35.5 g of chlorine gas C. Atoms in 6 h graphite D. Atoms in 24 g of magnesium
10	Which of the following contains the same number of molecules as 9 g of water.	A. 2 g of hydrogen gas B. 14 g of nitrogen gas C. 32 g of oxygen gas D. 44 g of carbon dioxide gas
11	What has a mass equal to that of one mole of water.	A. 22.4 dm <sup>3</sup> of water B. One mole of steam C. One molecule of water D. Two moles of hydrogen molecules and one mole of oxygen molecules.
12	$12.04 \times 10^{23}$ atoms of nitrogen gas is equal to.	A. 1 mol B. 2 mol C. 3 mol D. 4 mol
13	Which statement is incorrect about 64 g of SO <sub>2</sub> .	A. It is one mole SO <sub>2</sub> B. The number of SO <sub>2</sub> molecule are $6.02 \times 10^{23}$ C. The number of oxygen atoms are $6.02 \times 10^{23}$ D. The number sulphur atom are $6.02 \times 10^{23}$

14	How many atoms are present in half mol of oxygen gas. Gas exist in diatomic state.	B. $6.02 \times 10^{23}$ C. $2 \times 10^{23}$ D. $1.003 \times 10^{23}$
15	How many moles of CO are present having $12.04 \times 10^{23}$ molecules of CO <sub>2</sub> .	A. 0.5 mol B. 1.0 mol C. 1,5 mol D. 2.0 mol
16	A glass is full of water and contains $6.02 \times 10^{23}$ molecules of H <sub>2</sub> O The mass of water molecules is.	A. 18 gm B. 90 g C. 120 g D. 180 g
17	The mass fo $1.505 \times 10^{23}$ atoms of sulphur is.	A. 0.5 g B. 0.6 g C. 0.7 g D. 0.8 g