

## Chemistry 9th Class English Medium Unit 6 Online Test

Sr	Questions	Answers Choice
1	The characteristics of reversible reactions are the following except.	A. Product never recombine to form reactants B. They never complete C. They have a double arrow between reactants and products D. They proceed in both ways
2	In an irreversible reaction equilibrium	A. Never established B. Established quickly C. Established slowly D. Established when reaction stops
3	What will happen to the concentrations of the product if a reversible reaction at equilibrium is not disturbed.	A. They will keep on increasing B. They will keep on decreasing C. They will remain constant D. They will remain constant for some time and then start decreasing
4	If reversible reaction useful for preparing compounds on large scale.	A. Yes B. No C. They are useful only when equilibrium lies far to the left side D. They are useful only when equilibrium lies far to the right side
5	When a reaction will become a reversible one?	A. If the activation energy of the forward reaction is comparable to that of backward reaction B. If the activation energy of the forward reaction is higher than that of backward reaction C. If the activation energy of the forward reaction is lower than that of backward reaction D. If the enthalpy change of both the reactions is zero.
6	In an irreversible reaction equilibrium is.	A. The forward reaction will be favoured B. No effect on forward or backward reaction C. No effect on backward reaction D. The backward reaction will be favoured
7	Why the gas starts coming out when you open a can of fizzy drink.	A. Because the solubility of the gas increases B. Because the gas is dissolved under pressure hence it comes out when pressure is decreased C. Because the gas is insoluble in water D. Because the solubility of the gas decreases at high pressure.
8	What condition should be met for the reversible reaction to achieve the state of equilibrium.	A. The concentration of all the reactants and the product should become constant B. All the reactants should be converted into the product C. 50% of the reactant should be converted into products. D. One of the products should be removed from the reaction mixture.
9	CaO or lime is used extensively in steel, glass and paper industries. It is produced in an exothermic reversible reaction by the decomposition of $\text{CaCO}_3$ . Choose the conditions to produce maximum amount of lime.	A. Heating at high temperature in an open vessel B. Heating at high temperature in a closed vessel C. Cooling it in a closed vessel D. Cooling it in an open vessel
	An inorganic chemist places one mole of $\text{PCl}_5$ in container A and one mole of each $\text{Cl}_2$ and	A. Both the containers will have zero concentration of its reactants. B. Both the containers will have the

10	<p>PCl<sub>3</sub> in container B. Both the containers were sealed and heated to the same temperature to reach the state of equilibrium. Guess about the composition of mixtures in both the containers.</p>	<p>same composition of mixtures            C. Container A will have more concentration of PCl<sub>3</sub> than B.            D. Container A will have less concentration of PCl<sub>3</sub> than B.</p>
11	<p>Predict which components of the atmosphere react in the presence of lightning.</p>	<p>A. N<sub>2</sub> and H<sub>2</sub>O            B. O<sub>2</sub> and H<sub>2</sub>O            C. N<sub>2</sub> and O<sub>2</sub>            D. CO<sub>2</sub> and O<sub>2</sub></p>
12	<p>What will happen if the rates of forward and reverse reactions are very high</p>	<p>A. The reaction will be practically irreversible            B. The equilibrium point will reach very soon            C. The equilibrium point will reach very late            D. The reaction will not attain the state of dynamic equilibrium</p>