

Chemistry 10th Class English Medium Unit 10 Online Test

Sr Questions Answers Choice 1 Citric acid is present in A Citrus fulls B. Sour milk C. Rancid butter D. Apple 2 Which acid used for food preservation? A H-sub/2e/slub-SO-sub/3e/slub-COH 3 Which acid is used for etching designs on copper plates? A H-sub/2e/slub-SO-sub/3e/slub-COH 4 Which acid is used an electrolyte in lead storage battery? A H-sub/2e/slub-SO-sub/3e/slub-COH 5 Which one of the mineral acid. A H-sub/2e/slub-SO-sub/3e/slub-COH 6 When acid reacts with sulphites and Bi sulphates which gas is evolved? A H2 B. CO-sub/2e/slub-SO-sub/3e/slub-COH 7 When acid react with carbonates and bicarbonates which gas is evolved? A H2 B. CO-sub/2e/slub-COH B. C			
1 Citric acid is present in B. Sour milk C. Rancid butter D. Apple 2 Which acid used for food preservation? A Hesub>2 2 Which acid used for food preservation? B. HOVSub>3 3 Which acid is used for etching designs on copper plates? A Hesub>2 4 Hesub>2 3 Which acid is used an electrolyte in lead storage battery? B. HOVSub>3 4 Hesub>2 4 Which acid is used an electrolyte in lead storage battery? B. HOVSub>3 5 HOVSub>3 5 Which one of the mineral acid. A HEsub>2 5 Which one of the mineral acid. A HZ B. HSub>2 6 When acid reacts with sulphites and Bi sulphates which gas is evolved? B. C Sub>2 7 When acid react with carbonates and bicarbonates which gas is evolved? B. C Sub>2 8 When acids react with metals which gas is evolved? B. C Sub>2 8 When acids react with metals which gas is evolved? B. C Sub>2 9 Which one is Lewis Acid? B. HSub>2 9 Which one is Lewis Acid? B. A Salt B. Water C. Addiuct C. C Sub>3 9 A product of any Lewis acid base reaction is a single specie called. B. B. Water C. Addiuct D. None of these 11 According to Lewis concept a base is a substance which can donate. B. Base C. Amphoteric	Sr	Questions	Answers Choice
2 Which acid used for food preservation? C. HOL C. H	1	Citric acid is present in	B. Sour milk C. Rancid butter
Sh. HNO-sub-3-/sub> C. HCI D. CH-sub-3-/sub> C. HCI D. CH-sub-3-/sub-4-/sub-3-/sub-4-/sub-3-/sub-4-/sub-3-/sub-4-/sub-3-/sub-4-/sub-3-/sub-4-/sub-3-/sub-3-/sub-4-/sub-3-/sub-3-/sub-4-/sub-3-/sub-3-/sub-3-/sub-4-/sub-3	2	Which acid used for food preservation?	B. HNO ₃ C. HCl
4 Which acid is used an electrolyte in lead storage battery? B. HNO'sub>3-/sub> C. HCI D. CH'sub>3-/sub>COOH A HCI B. Hssub>2-/sub>SO(sub>4-/sub> C. HO) and these Mich one of the mineral acid. A HCI B. Hssub>2-/sub>SO(sub>4-/sub> D. All of these A HCI B. Hssub>3-/sub> D. All of these A HCI B. HSsub>3-/sub> D. All of these A HCI B. HSsub>3-/sub> D. All of these A HCI B. CO sub>2-/sub> D. NH-sub>3-/sub> D. All of these A HCI B. CO sub>2-/sub> D. NH-sub>3-/sub> D. NH-sub>3-/sub> D. NH-sub>3-/sub> D. NH-sub>3-/sub> D. NH-sub>2-/sub> D. NH-sub-2-/sub> D. NH-sub-2-/sub-2	3	Which acid is used for etching designs on copper plates?	B. HNO ₃ C. HCI
5 Which one of the mineral acid. B. H*sub>2 C. HNO ₃ C. HNO ₃ D. All of these A. H2 B. CO ₂ C. SO ₂ C. SO ₂ D. NH*sub>3 D. NH*sub>3 D. NH*sub>3 A. H2 B. CO ₂ C. SO ₂ D. NH*sub>3 A. H2 B. CO ₂ C. CI ₂ D. NH*sub>3 A. H2 B. CO ₂ C. CI ₂ C. CI ₂ D. Nsub>2 D. All of these A. BF ₃ D. All of these A. Salt B. Water C. FeC(rsub>3 D. All of these A. Salt B. Water C. Adduct D. None of these A. Proton B. Electron pair C. Neutron D. Electron D. Electr	4	Which acid is used an electrolyte in lead storage battery?	B. HNO ₃ C. HCI
8 CO-sub>2 C SO(sub>2 C SO(sub)>2 C SO(sub)>2 D NI+sub>3 A H2 B CO-sub>2 C SO(sub)>2 C SO(sub)>2 D NI+sub>3 A H2 B CO(sub)>2 C C (I+sub>2 C C (I+sub>2 D NS-sub>2 D NS-sub>2 D NS-sub>2 D NS-sub>2 A H2 B CO(sub)>2 D NS-sub>2 D NS-sub>2 C C (I+sub)>2 D NS-sub>2 D NI-sub>2 C FeCI-sub>3 C FeCI-sub>3 D NI-sub-2 D NI-sub-2 D NI-sub-2 D NI-sub-2 D NI-sub-2 D NI-sub-2 C FeCI-sub>3 C FeCI-sub>3 C NI-sub-2 D NI-sub-2 D NI-sub-2 D NI-sub-2 D NI-sub-2 D NI-sub-2 D NI-sub-2 C FeCI-sub>3 C FeCI-sub>3 C NI-sub-2 D NI-sub-2 <td>5</td> <td>Which one of the mineral acid.</td> <td>B. H₂SO₄C. HNO₃</td>	5	Which one of the mineral acid.	B. H ₂ SO ₄ C. HNO ₃
When acid react with carbonates and bicarbonates which gas is evolved? B. CO ₂ C. CI ₂ D. N ₂ A. H ₂ B. O ₂ B. O ₂ C. CI ₂ A. H ₂ C. CI ₂ D. N ₂ C. CI ₂ C. CI ₂ D. N ₂ C. CI ₂ D. N ₃ C. CI ₃ D. All of these A. Salt B. Water C. Adduct D. None of these A. Proton B. Electron pair C. Neutron D. Electron C. Approach D. A Acid B. Base C. Amphoteric	6	When acid reacts with sulphites and Bi sulphates which gas is evolved?	B. CO ₂ C. SO ₂
8 When acids react with metals which gas is evolved? 9 Which one is Lewis Acid? A BF ₂ D. N ₂ A. BF ₃ B. AICI ₃ C. FeCI ₃ D. All of these A Salt B. Water C. Adduct D. None of these 11 According to Lewis concept a base is a substance which can donate. A Substance which can behave as an acid as well as a base is called. A Substance which can behave as an acid as well as a base is called. A Proton B. Electron pair C. Neutron D. Electron A. Acid B. Base C. Amphoteric	7	When acid react with carbonates and bicarbonates which gas is evolved?	B. CO ₂ C. Cl ₂
9 Which one is Lewis Acid? 10 A product of any Lewis acid base reaction is a single specie called. A product of any Lewis acid base reaction is a single specie called. A Salt B. Water C. Adduct D. None of these A. Proton B. Electron pair C. Neutron D. Electron D. Electron A. Acid B. Base C. Amphoteric	8	When acids react with metals which gas is evolved?	B. O ₂ C. Cl ₂
A product of any Lewis acid base reaction is a single specie called. B. Water C. Adduct D. None of these A. Proton B. Electron pair C. Neutron D. Electron A. Acid B. Base C. Amphoteric	9	Which one is Lewis Acid?	B. AlCl ₃ C. FeCl ₃
According to Lewis concept a base is a substance which can donate. B. Electron pair C. Neutron D. Electron A. Acid B. Base C. Amphoteric	10	A product of any Lewis acid base reaction is a single specie called.	B. Water C. Adduct
12 A substance which can behave as an acid as well as a base is called. B. Base C. Amphoteric	11	According to Lewis concept a base is a substance which can donate.	B. Electron pair C. Neutron
	12	A substance which can behave as an acid as well as a base is called.	B. Base C. Amphoteric