

ECAT Pre General Science Online Test

Sr	Questions	Answers Choice
1	How many isotopes of helium are present?	A. 1 B. 2 C. 3 D. 4
2	The number of isotopes of hydrogen are	A. 2 B. 1 C. 3 D. 4
3	Nuclei that have the same charge number but different mass number are called	A. isotones B. isomers C. isotopes D. isobars
4	Electrons are	A. positive charged B. negatively charged C. massless D. neutral
5	Neutrons are	A. positive charge B. negatively charged C. massless D. neutral
6	The diameter of an atom is of the order	A. 10^{-125} m B. 10^{-11} m C. 10^{-10} m D. 10^{-9} m
7	Structure of the nucleus was explained by	A. J.J Thomson B. Bohr C. Millikan D. Rutherford
8	Charge on proton is	A. 1.59×10^{-9} C B. 1.59×10^{-7} C C. -1.59×10^{-19} C D. 1.59×10^{-19} C
9	Mass of proton is of order of	A. 10^{-31} gm B. 10^{-27} kg C. 10^{-24} gm D. 10^{+27} kg
10	The number of neutrons in the nucleus of ${}_{92}\text{U}^{235}$ are	A. Infinite B. 92 C. 235 D. 143
11	For an atom having atomic number Z and atomic weight A, the number of electron in an atoms	A. A - Z B. A + Z C. Z D. A
12	For an atom having atomic number Z and atomic weight A, the charge on the nucleus is	A. A - Z B. A + Z C. Z D. A
13	The number of all the protons and neutrons in a nucleus is known as	A. atomic number B. mass number C. charge number D. none of these
14	The number of protons inside a nucleus is called	A. mass number B. atomic weight C. atomic number D. none of these
15	The total charge of any nucleus is given as	A. Ze^{+2} B. $Z^{+2}e$ C. Ze D. Ze

16	The nucleus of uranium -235 differs from a nucleus of a uranium -238 in that the latter contains	A. 3 more neutrons B. 3 more electrons C. 3 more protons D. 3 more ions
17	For an atom having atomic number 'Z' and atomic weight 'A', the number of neutrons in the nucleus is	A. $A - Z$ B. A C. Z D. $A + Z$
18	According to Rutherford atomic model, the positive charge in an atom	A. is concentrated at its centre B. is in the form of positive electron at same distance from its centre C. is spread uniformly through its volume D. none of these
19	The chemical behaviour of an atom is determined by	A. binding energy B. atomic number C. mass number D. number of isotopes
20	1 amu is equal to	A. 1.66×10^{-24} kg B. 1.66×10^{-19} kg C. 1.66×10^{-34} kg D. 1.66×10^{-27} kg
21	Mass of proton is	A. 1.67×10^{-27} kg B. 1.67×10^{-31} kg C. 1.66×10^{-34} kg D. 1.67×10^{-17} kg
22	Mass of neutron is	A. 1.67×10^{-31} kg B. 1.67×10^{-27} kg C. 9.1×10^{-31} kg D. 1.67×10^{-19} kg
23	Nucleus consists of	A. proton and neutron B. protons and electron C. electron and neutron D. protons only
24	A particle having the mass of electron and charge of a proton is called a	A. photon B. positron C. antiproton D. antineutrino
25	Charge on neutron is	A. 1.6×10^{-19} C B. zero C. -1.6×10^{-19} C D. 1.2×10^{-19} C
26	In 1932 Chadwick discovered	A. proton B. neutron C. photon D. electron
27	Neutron was discovered by	A. Curie B. Roentgen C. Chadwick D. Rutherford
28	Neutron was discovered in	A. 1915 B. 1920 C. 1925 D. 1932
29	Proton was discovered by Rutherford in	A. 1915 B. 1906 C. 1910 D. 1920
30	According to the electromagnetic wave theory of light, increasing the intensity of incident light should increase the	A. number of photoelectrons B. size of the photoelectrons C. charge on photoelectrons D. K.E of photoelectrons