ECAT Pre General Science Physics Chapter 21 Nuclear Physics Online Test

Sr	Questions	Answers Choice
1	Hydrogen atom with only one proton and one neutron in its nucleus, and one electron, is called	A. deuterium B. protium C. tritium D. none of these
2	Hydrogen atom with only one proton in its nucleus, and one electron in its orbit is called	A. deuteron B. deterium C. protium D. tritium
3	How many isotopes of helium are present?	A. 1 B. 2 C. 3 D. 4
4	The number of isotopes of hydrogen are	A. 2 B. 1 C. 3 D. 4
5	Nuclei that have the same charge number but different mass number are called	A. isotones B. isomers C. isotopes D. isobars
6	Electrons are	A. positive charged B. negatively charged C. massless D. neutral
7	Neutrons are	A. positive charge B. negatively charged C. massless D. neutral
8	The diameter of an atom is of the order	A. 10⁻¹²⁵m B. 10⁻¹¹m C. 10⁻¹⁰m D. 10⁻⁹m
9	Structure of the nucleus was explained by	A. J.J Thomson B. Bohr C. Millikan D. Rutherford
10	Charge on proton is	A. 1.59 x 10⁻⁹C B. 1.59 x 10⁻⁷C C. -1.59 x 10⁻¹⁹C D. 1.59 x 10⁻¹⁹C
11	Mass of proton is of order of	A. 10⁻³¹gm B. 10⁻²⁷kg C. 10⁻²⁴gm D. 10⁺²⁷kg
12	The number if neutrons in the nucleus of $_{92}U^{235}$ are	A. Infinite B. 92 C. 235 D. 143
13	For an atom having atomic number Z and atomic weight A, the number of electron in an atoms	A. A - Z B. A + Z C. Z D. A
14	For an atom having atomic number Z and atomic weight A, the charge on the nucleus is	A. A - Z B. A + Z C. Z D. A
15	The number of all the protons and neutrons in a nucleus is known as	A. atomic number B. mass number C. charge number D. none of these

16	The number of protons inside a nucleus is called	A. mass number B. atomic weight C. atomic number D. none of these
17	The total charge of any nucleus is given as	A. Ze² B. Z²e C. Z/e D. Ze
18	The nucleous of uranium -235 differs from a nucleous of a uranium -238 in that the later contains	A. 3 more neutrons B. 3 more electrons C. 3 more protons D. 3 more ions
19	For an atom having atomic number 'Z' and atomic weight 'A', the number of neutrons in the nucleous is	A. A - Z B. A C. Z D. A + Z
20	According to Rutherford atomic model, the positive charge in an atom	A. is concentrated at its centre B. is in the form of positive electron at same distance from its centre C. is spread uniformly through its volume D. none of these
21	The chemical behaviour of an atom is determined by	A. binding energy B. atomic number C. mass number D. number of isotopes
22	1 amu is equal to	A. 1.66 x 10⁻²⁴kg B. 1.66 x 10⁻¹⁹kg C. 1.66 x 10⁻³⁴kg D. 1.66 x 10⁻²⁷kg
23	Mass of proton is	A. 1.67 x 10⁻²⁷kg B. 1.67 x 10⁻³¹kg C. 1.66 x 10⁻³⁴kg D. 1.67 x 10⁻¹⁷kg
24	Mass of neutron is	A. 1.67 x 10⁻³¹kg B. 1.67 x 10⁻²⁷kg C. 9.1 x 10⁻³¹kg D. 1.67 x 10⁻19kg
25	Nucleus consists of	A. proton and neutron B. protons and electron C. electron and neutron D. protons only
26	A particle having the mass of electron and charge of a proton is called a	A. photon B. position C. antiproton D. antineutrino
27	Charge on neutron is	A. 1.6 x 10⁻¹⁹C B. zero C. -1.6 x 10⁻¹⁹C D. 1.2 x 10⁻¹⁹C
28	In 1932 Chadwick discovered	A. proton B. neutron C. photon D. electron
29	Neutron was discovered by	A. Curie B. Roentgen C. Chadwick D. Rutherford
30	Neutron was discovered in	A. 1915 B. 1920 C. 1925 D. 1932