

ECAT Physics Chapter 11 Heat & Thermodynamics Online Test

Sr	Questions	Answers Choice
1	The value of E_{in} in coulomb's law is:	A. $9 \times 10^9 \text{ Nm}^2 \text{ C}^{-2}$ B. $8.85 \times 10^{12} \text{ C}^2 \text{ N}^{-1} \text{ m}^{-2}$ C. $8.85 \times 10^{12} \text{ Nm}^2 \text{ C}^{-2}$ D. $9 \times 10^9 \text{ C}^2 \text{ N}^{-1} \text{ m}^{-2}$
2	If the distance between two charges is doubled, the force between them will become:	A. Double B. Half C. Three times D. One fourth E. One third
3	Average KE of a gas molecule has:	A. Direct relation with absolute temperature and inverse relation with pressure B. Direction relation with both absolute temperature and pressure C. Inverse relation with both absolute temperature and pressure D. None of these
4	Gas constant per molecule is called:	A. Universal gas constant B. Stefan's constant C. Boltzmann constant D. Gravitation constant
5	Tick the correct pair when M denotes the molecular mass and other symbols carry usual meanings:	A. $N = nN_A$, $M = MN_A$ B. $n = N/N_A$, $M = mN_A$ C. $M = N_A/N$, $N = m/n$ D. $N = nN_A$, $M = mN_A$
6	In the formula $P = N_0KT$, N_0 denotes:	A. Number of molecules per unit per volume B. Number of moles C. Number of molecules D. None of these
7	The value of universal gas constant R is:	A. 8.314 J/K mole K B. 8314 J/K mole K C. 8.314 J/mole K D. None of these
8	If the formula $PV = nRT$, n denotes:	A. Number of molecules per unit volume B. Number of moles C. Number of molecules D. None of these
9	While deriving equation of pressure by kinetic theory of gases, we take into account:	A. Only linear motion of molecules B. Only rotational motion C. Only vibratory motion D. All of these
10	Pressure applied at any point of gas at rest is transmitted equally to all parts of the gas. This is the statement of:	A. Newton's second law B. Pascal's law C. Carnot theorem D. Second law of thermodynamics
11	If n denotes the total number of molecules in cubic vessel such that m is mass of each molecule and l is length of each side of vessel, then nm/l^3 gives the:	A. Force B. Density C. Work done D. Pressure
12	The rate of change of momentum of a molecule is equal to:	A. Pressure B. Work C. Force D. None of these

C. Density
D. Force

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If a molecule with momentum mv strikes a wall and rebound then the change in momentum will be:

A. $-2\ mv$
B. Zero
C. $2\ mv$
D. mv

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Pressure may be define as _____ per second per unit area:

A. Change in force
B. Change in momentum
C. Change in energy
D. Work done

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Truth of kinetic energy is confirmed by:

A. Diffusion of gases
B. Brownian motion
C. Both A and B
D. None of these