ECAT Chemistry Online Test

Sr	Questions	Answers Choice
1	Hemoglobin contains nearly:	A. 10,000 atoms B. 100 atoms C. 1000 atoms D. 1 atom
2	Hemoglobin is 68000 times heavier than:	A. Oxygen atom B. Nitrogen atom C. Carbon atom D. Hydrogen atom
3	NH_3, HCL, H_2O, HL are:	A. <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Diatomic molecules<o:p></o:p> B. <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Poly-atomic molecules<o:p></o:p> C. <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Mono-atomic molecules<o:p></o:p> D. <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Hetero atomic molecules<o:p></o:p>
4	He Ar and Ne are:	A. <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Mono-atomic molecules B. <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Hetero atomic molecules<o:p></o:p> C. <span style="font-size: 10.5pt; line-height:

		107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Poly-atomic molecules<o:p></o:p> D. <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Diatomic molecules<o:p></o:p>
5	CL_2, N_2 and O_2 are:	A. <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Diatomic molecules<o:p></o:p> B. <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Hetero atomic molecules<o:p></o:p> C. <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Poly-atomic molecules<o:p></o:p> D. <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Mono-atomic molecules
6	$C_6H_{12}O_6$and $C_{12}H_{22}O_{11}$ are:	A. Mono-atomic molecules B. Diatomic molecules C. Poly-atomic molecules D. <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Hetero atomic molecules<o:p></o:p>
7	The number of subatomic particles in atoms sidcovered is more than:	A. 110 B. 100 C. 125 D. 90
8	Atoms can be evident by use of electron microscope, field ionization microscope and:	A. x-rays B. Video camera C. Telescope D. Compound microscope
9	Determination of atomic masses and invention of system of writing symbols was made by:	A. J. Berzelius B. Democritus C. Dalton D. None of above

10	First atomic theory was put forward by an English school teacher:	A. Maxewell B. Newton C. Sanger D. John Dalton
11	Metal tend to lose electrons, becoming:	A. Metals<o:p></o:p> B. Positively charged<o:p></o:p> C. Negatively charged<o:p></o:p> D. (a) <!--[endif]-->And (c)<o:p></o:p>
12	Atoms and molecules can either gain or lose electrons, forming charge particles called:	A. Positrons<o:p></o:p> B. Photons<o:p></o:p> C. Ions<o:p></o:p> D. Electrons<o:p></o:p>
13	Covalent compound s mostly exist in the form of:	A. Protons<o:p></o:p> B. Atoms<o:p></o:p> C. <span style="font-size:11.0pt;line-height:107%; font-family:"Calibri","sans-serif";mso-ascii-theme-font:minor-latin;mso-fareast-font-family: Calibri;mso-fareast-theme-font:minor-latin;mso-hansi-theme-font:minor-latin; mso-bidi-font-family:"Times New Roman";mso-bidi-theme-font:minor-bidi; mso-ansi-language:EN-US;mso-fareast-language:EN-US;mso-bidi-language:AR-SA">Neutrons D. Molecules<o:p></o:p>
14	The diameter of atoms is of the order:	A. 2* 10⁻⁵m B. 2* 10⁻¹⁰m^{<o:p></o:p>} C. 2* 10⁻²m D. 2* 10⁻³m
15	CO+ is an example of	A. Stable molecule B. Anionic molecule ion C. Cationic molecular ion D. Free radical
16	When an electron is added to a uni positive ion we get:	A. Cation B. Molecule C. Neutral atom D. Anion
17	The number of atoms present in molecule determines its:	A. Molecularity B. Atomicity C. Basicity D. Acidity
18	Matter is defined as any thing which occupies space and:	A. Molecules B. Mass C. Compound D. Molecules
19	Smallest particle of an element which may or may not have independent existence is known	A. A molecule B. An ion

	as:	C. An atom D. An electron
20	The branch of science dealing with structure, composition and changes in matter and laws and principles which govern these changes is called as	A. chesmistry B. Geology C. Physics D. Mechanics
21	Isotopes differ in	A. properties which depend upon mass B. arrangement of electrons in orbitals C. chemical properties D. the extent to which they may be affected in electric fields
22	The tip of the funnel should touch the side of the beaker in order to avoid	A. Splashing B. Leakage C. Mixing D. Contamination
23	The tip of funnel should be 1 or 2cm larger than the circle of the	A. Beaker B. Solid C. Filter paper D. Liquid
24	The filtration process is used to separate solid from	A. Liquid B. Gas C. Solid D. All of above
25	The solid remained on filter paper during filtration is called the	A. Substance B. Residue C. Undue D. Filtrate
26	The sample being analyzed is called	A. Electrolyte B. Substance C. Analyte D. All of above
27	The detection of functional group is called	A. Numerical analysis B. Qualitative analysis C. Combustion analysis D. Quantitative analysis
28	Identification of a substance, determination of its structure and quantitative analysis of its composition are the aspects covered by	A. Modern analytical physics B. Mechanical chemistry C. Biochemistry D. Modern analytical chemistry
29	Estimation of Na in sea water is an example of	A. Numerical analysis B. Qualitative analysis C. Quantitative analysis D. None of above
30	A complete chemical characterization of a compound must include	A. Qualitative analysis B. Chemical analysis C. Quantitative analysis D. Both a and c