ECAT Chemistry Online Test

Sr	Questions	Answers Choice
1	Which of the following discoveries resulted in a version of the Mendeleefs periodic law	A. The nucleus of atom by Rutherford B. The elements polonium and radium by the Curies C. Atomic numbers by Moseley D. x-rays by Roentgen
2	The hydration energy is the heat evolved when one mole of gaseous ion is dissolved in water. The hydration energy of an ion	A. Increases with increase of charge to mass ratio B. Decreases with increase of charge to mass ratio C. Depends on sign of charge +ve or -ve D. Depends upon the solvent
3	Although hydrogen resemble with the elements of group IA, IVA and VII but it is usually placed in	A. Group IA B. Group IV A C. Group VII D. Group VIII
4	Which of the following is a member of -block	A. Zn B. Al C. B D. Br
5	The fourteen elements following actinium are known as	A. Lanthanones B. Lanthanides C. Rare earths D. Actinides
6	The fourth period contains elements	A. 8 B. 16 C. 18 D. 32
7	What is the nature of Al_2O_3	A. Acidic B. Basic C. Amphoteric D. Neutral
8	What is the nature of SO_2	A. Basic B. Strongly acidic C. Weakly acidic D. Amphoteric
9	Na_2O is	A. Acidic B. Basic C. Neutral D. Amphoteric
10	A pair of elements in the same family in the periodic classification is	A. Cl and C B. Ca and Al C. N and Ne D. Na and K
11	The elements of group IA are called	A. Chalocogens B. Halogens C. Alkali metals D. Alkaline earth metals
12	The oxides of electronegative elements are	A. Basic B. neutral C. Acidic D. Amphoteric
13	The coinage metals are	A. Ni, Pd, Pt B. Cu, Ag, Au C. Zn, Al, Pb D. Fe, Si, Sn
14	In a period, melting points of elements	A. Increases B. Decreases C. Remain constant D. First increases then decreases

15	Ionization energy depends upon	A. Nuclear charge B. Atomic size C. Shielding effect D. All of the above
16	In a group, the ionization energy	A. Increase B. Decreases C. Remain constant D. First increases then decreases
17	The amount of energy required to remove an electron from an atom of an element in the gaseous state is called	A. Electron affinity B. Electronegatively C. Ionization energy D. None of these
18	Which is the transition element among the following	A. B B. Al C. Cu D. Cs
19	In sixth period 14 of its transition elements are called	A. Lanthanides B. Actinides C. Radioactive elements D. None
20	The correct order of 2nd I.P. of C,N,O and F is	A. O > F > N > C B. O > N > F > C C. C > N > O > F D. F > O > N > C
21	Which is the longest period of the periodic table	A. 5th B. 7th C. 6th D. 2nd
22	Which orbital is in the process of completion in case of transition elements	A. p-orbital B. f-orbital C. d-orbital D. s-orbital
23	Number of elements in the first period of the periodic table are	A. Two B. Four C. One D. Eight
24	The energy absorbed when an electron is added to a gaseous atom to form a gaseous ion is called	A. Electron affinity B. Ionization energy C. Both of these D. None of these
25	The positive ion is always smaller than the neutral atom while the negative ion is always bigger than the neutral atom. The atomic and ionic radii of Na, F, Na^+, $F^-$are in ppm	A. Na F Na⁺ F⁻ 157 72 95 136 B. Na F Na⁺ F⁻ 157 95 172 136 C. Na F Na⁺ F⁻ 72 95 136 157 D. Na F Na⁺ F⁻ 157 136 95 72
26	The electropositive elements from	A. Acidic oxides B. Basic oxides C. Neutral oxides D. None
27	According to Mendleev, the physical and the chemical properties are the periodic function of their	A. Atomic number B. Atomic mass C. Atomic wt D. None
28	Transition elements have valence electrons in	A. s-orbital B. p-orbital C. d-orbital D. f-orbital
29	Rare earth elements are	A. s-block elements B. p-block elements C. d-block elements D. f-block elements
30	The IA elements are called	A. Alkaline earth metal B. Alkaline metals C. The halogens D. The inert gases