

ECAT Chemistry Online Test

Sr	Questions	Answers Choice
1	Which of the following is not the reactions of lithium	
2	Gypsum is a common mineral of	A. Magnesium B. Strontium C. Calcium D. Barium
3	Borax is a common mineral of alkali metal sodium. Its formula is	A. $\text{Na}_2\text{B}_4\text{O}_7$ B. $\text{Na}_2\text{B}_4\text{O}_7 \cdot 10\text{H}_2\text{O}$ C. $\text{Na}_2\text{B}_3\text{O}_6 \cdot 10\text{H}_2\text{O}$ D. $\text{Na}_2\text{B}_4\text{O}_7 \cdot 5\text{H}_2\text{O}$
4	Which of the following is lighter	A. Li B. k C. Na D. Ca
5	The alkaline earth elements have in their s-orbital	A. One electron B. Two electron C. No electron D. Three electron
6	Which of the element is not an alkaline earth metal	A. Beryllium B. Strontium C. Barium D. Caesium
7	Which of the element is not alkali metal	A. Lithium B. Rubidium C. Francium D. Magnesium
8	From left to right, atomic radii of transition elements	A. Increases B. Decreases C. Remain same D. None of the above
9	The coinage metals are	A. Ni, Pd, Pt B. Cu, Ag, Au C. Zn, Al, Pb D. Fe, Si, Sn
10	Which is the most volatile compound	A. HI B. HCl C. HBr D. HF
11	Mark the correct statement	A. Na^+ is smaller than Na atom B. Na^+ is larger than Na atom C. Cl^- is the smaller than Cl atom D. Cl^- (ion) and Cl (atom) are equal in size
12	Keeping in view the size of atoms, which order is the correct one	A. Mg > Sr B. Ba > Mg C. Lu > Ce D. Cl > I
13	The valency, ionization energy and electronegativity of elements are related to its	A. Atomic number B. Properties C. Atomic weight D. Family group
14	Which of the following has highest oxidation potential	A. Be B. Li C. Na D. Ca

15	The valence shell of hydrogen is half filled like those of	A. IV - A B. VIA C. V - A D. VIIA
16	Which of the following statement about electron affinity of two elements is correct	A. Carbon has greater than oxygen B. Sulphur has less than oxygen C. Iodine has greater than bromine D. Bromine has less than chlorine
17	The structure of complex hydrides is	A. Tetrahedral B. Trigonal C. Octahedral D. Square planar
18	For the representative elements from left to right across a period in the periodic table, the electron affinity of the atom generally	A. Increases B. Remains constant C. Decreases D. Not clear
19	The elements of f-block are also known as	A. Inner-transition B. Outer transition C. Normal elements D. Alkaline earth metals
20	Elements in the same family have	A. Same atomic number B. Molecular wt same C. Same chemical properties D. Same electronic configuration
21	According to the periodic law, the chemical properties of the elements are periodic functions of their	A. Density B. Atomic number C. Atomic mass D. Mass number
22	Which is not interstitial hydride	A. LaH B. VH C. TaH D. None
23	In a group from top to bottom, the hardness of alkali metals	A. Remains unchanged B. Increases C. Decreases D. None
24	Which has highest 1st I.E.	A. Br B. Cl C. F D. I
25	Each vertical column of the periodic table includes elements with chemical characteristics that are in general	A. Identical B. Similar C. Different D. Similar as well as different
26	Which element should have the greatest value for electronegativity when combined with hydrogen	A. Na B. Si C. S D. Cl
27	Which of the following elements have the largest radius	A. F B. Cl C. Br D. I
28	Which of the following elements should be the least metallic in character	A. Rb B. In C. Te D. I
29	Gradation in properties in the periods of periodic tables are due to change in	A. Atomic weight B. The number of electrons C. Number of protons D. Electronic configuration
30	The most distinctive character among the elements is their division into	A. Metals and non-metals B. Solids, liquids and gases C. Atoms and molecules D. Active and inactive elements