

## ECAT Chemistry Online Test

Sr	Questions	Answers Choice
1	Which of the following isoelectronic ions has the lowest ionization energy?	A. $K^{+}$ B. $Ca^{2+}$ C. $Cl^{-}$ D. $S^{2-}$
2	The atomic radius increases as we move down a group because	A. Effective nuclear charge increases B. Atomic mass increases C. Additive electrons are accommodated in new electron level D. Atomic number increases
3	Which of the following statement about fluorine is not correct?	A. Electron affinity of chlorine is greater than that of fluorine B. Bond energy of fluorine is less than that of chlorine C. Fluorine cannot be prepared by electrolysis of fused metal fluorides D. Fluorine does not form oxoacid
4	The element with highest electron affinity among the halogen is	A. F B. Cl C. Br D. I
5	The ionization potential is lowest for the	A. Halogens B. Inert gases C. Alkaline earth metals D. Alkali metals
6	In the modern long form of the periodic table elements are arranged in the increasing order of	A. Atomic mass B. Atomic number C. Mass number D. Isotopic number
7	Variable valency is characteristic of	A. Halogen B. Transition elements C. Alkali metals D. Noble gas
8	The correct order of electron affinity among the following is	A. $F > Cl > Br$ B. $Br > Cl > F$ C. $Cl > F > Br$ D. $F > Br > Cl$
9	Which of the following is most electronegative?	A. Carbon B. Silicon C. Lead D. Tin
10	Which of the following element has the maximum electron affinity?	A. F B. S C. I D. Cl
11	Which of the following has highest first ionization potential?	A. Carbon B. Oxygen C. Nitrogen D. Boron
12	Which of the following species has the highest ionization potential?	A. Ne B. $Al^{+}$ C. $Mg^{+}$ D. $Li^{+}$
13	Gradual addition of electronic shells in the noble gases causes a decrease in their	A. Ionization energy B. Atomic radius C. Boiling point D. Density
14	In the periodic table, the element with atomic number 16 will be placed in the group	A. Fourteen B. Sixteen C. Thirteen D. Fifteen

15	Among the elements given below, the one with highest electropositivity is	A. Cu B. Cs C. Cr D. Ba
16	The correct arrangement of increasing order of atomic radius among Na, K, Mg, Rb is	A. Mg &lt; K &lt; Na &lt; Rb B. Mg &lt; Na &lt; K &lt; Rb C. Mg &lt; Na &lt; Rb &lt; K D. Na &lt; K &lt; Rb &lt; Mg
17	Alkali metals in each period have	A. Smallest size B. Lowest IE C. Highest IE D. Highest electronegativity
18	The element with atomic number 55 belongs to which block of the periodic table	A. s-block B. p-block C. d-block D. f-block
19	The attraction that an atom exerts on a pair of electrons that are being shared with another atom for forming covalent bond is referred to as its	A. Electron affinity B. Electronegativity C. Ionisation energy D. Valency
20	Among O, C, F, Cl, Br, the correct order of increasing radii is	A. F O C Cl Br B. F C O Cl Br C. F Cl Br O C D. C O F Cl Br
21	Among the following elements which one has the highest value of first ionization potential?	A. Oxygen B. Argon C. Barium D. Cesium
22	The element with atomic number 26 will be found in group	A. 2 B. 8 C. 6 D. 10
23	Which of the following pair of atomic numbers represents s-block elements?	A. 7, 15 B. 6, 12 C. 9, 17 D. 3, 20
24	The valence shell electronic structure of an element is $ns^2np^5$ . The element will belong to the group of	A. Alkali metals B. Inert metals C. Noble gases D. Halogen
25	The valency of noble gases, in general, is	A. Zero B. One C. Three D. Two
26	Which among the following elements have lowest value of $IE_1$ ?	A. Pb B. Sn C. Si D. C
27	Of the given alkali metals, the one with smallest size is	A. Rb B. Cs C. K D. Na
28	Which of the following elements is/are not liquid at 30°C?	A. Ga B. Hg C. Ge D. Cs
29	Which of the following metal requires radiation of highest frequency to cause emission of electrons?	A. Na B. Mg C. K D. Ca
30	Which of the following does not reflect the periodicity of elements?	A. Bonding behaviour B. Electronegativity C. Ionisation potential D. Neutral/proton ratio