ECAT Chemistry Chapter 9 Solutions Online Test

Sr	Questions	Answers Choice
1	What mass of NaOH is required to prepare 2.5 dm^3 of 1.5 M NaOH solution	A. 130 g B. 140 g C. 150 g D. 160 g
2	How many cm^3 of 1 M H_2SO_4 required to neutralize 10 cm^3 of 1 M NaOH	A. 2 cm^3 B. 2.5 cm^3 C. 5 cm^3 D. 10 cm^3
3	The term ebullioscopy is used for	A. Depression of freezing point B. Elevation in boiling point C. Lower of vapour pressure D. None of the above
4	What is the molarity of a solution containing 15.0 g urea in 500 cm^3 of solution	A. 0.5 M B. 1 M C. 1.5 M D. 2 M
5	Precipitation will occur until the ionic product becomes	A. Equal to K_{sp} B. Lesser than K_{sp} C. Greater than K_{sp} D. None of these
6	If ionic product of a solution is greater than solubility product, the solution is	A. Supersaturated B. Saturated C. Unsaturated D. None of these
7	If the ionic product of a solution is less than the solubility product, the solution is	A. Supersaturated B. Unsaturated C. Ideal D. Saturated
8	The process in which water molecules surround solute particles is called	A. Hydration B. Salvation C. Hydrolysis D. Dehydration
9	The process in which the solvent molecules are surrounded and interact with solute ions or molecules is called	A. Solvation B. Hydration C. Hydrogenation D. None
10	Solution may have units	A. Molarity B. Molality C. Mole fraction D. All of them
11	Number of moles of solute dissolved in 1 Kg of solvent is knowns as	A. Molarity B. Formality C. Molality D. Mole fraction
12	Number of moles of the solute dissolved per dm^3 of the solution is knows as	A. Molarity B. Formality C. %age D. None of these
13	A solution can be	A. Dilute and concentrated B. Saturated and dilute C. Saturated and unsaturated D. Supersaturated and saturated
14	A homogeneous mixture of two or more than two chemical substances is called	A. Solute B. Solution C. Solvent D. Salvation