ECAT Chemistry Chapter 7 Thermo Chemistry Online Test

Sr	Questions	Answers Choice
1	Work is a	A. State function B. Only function C. Non-state function D. State
2	When a system absorbs energy, the sign of delta E is	A. Neither positive nor negative B. Negative C. Positive D. None of above
3	When no work is done by the system	A. The volume of system decreases B. The volume of system increases C. The volume of system does not change D. None of above
4	By state, we mean the	A. Reaction of system B. Scope of a system C. Condition of a system D. None of above
5	Some non-spontaneous processes can be made to take place by supplying energy to the system from	A. Internal source B. Any source C. External source D. All of above
6	Burning of coal and hydrocarbon in air are examples of	A. Non-spontaneous reaction B. Spontaneous reaction C. Natural reaction D. Both b and c
7	A reaction will also be called a spontaneous if	A. It does not need energy to start with B. It needs energy to carry the whole process C. It needs energy at the end of reaction D. It needs energy to start with
8	The reaction of Zinc with copper sulphate solution is an example of	A. Oxidation reduction reaction B. Spontaneous reaction C. Spontaneous redox reaction D. Non-spontaneous reaction
9	When a piece of zinc is added to the copper sulphate solution, _____ colour of solution disappear	A. Pink B. Purple C. Blue D. Brown
10	Neutralization of a strong acid with a strong base is	A. Natural acid base reaction B. Artificial acid base reaction C. Spontaneous acid base reaction D. both a and c
11	Which one of the following is not related to spontaneous process	A. Unidirectional B. Real C. Irreversible D. Artificial
12	A process which takes place on its own without any outside assistance and moves from a non-equilibrium state towards an equilibrium state is termed as	A. Spontaneous process B. Natural process C. Non-spontaneous process D. Both a and b
13	The subject matter of first law of thermochemistry is based on	A. First law of Thermochemistry B. First law of Thermodynamics C. Second law of Thermochemistry D. Second law of Thermodynamics
14	The energy units in which heat changes usually expressed in SI-system are	A. Joule B. Calorie C. Kilo Joule D. Both a and c

15	In a chemical change, the energy in the form of heat will either be evolved or absorbed and this is called	A. Endothermic B. Heat of products C. Exothermic reaction D. Heat of reaction
16	It is noticed that energy in the form of heat is either evolved or absorbed as a result of a	A. Physical change B. Chemical change C. Biological change D. All of above
17	The majority of reactions which give stable products are	A. Exothermic B. Isothermal C. Endothermic D. Both a and c
18	In endothermic reactions, the heat contents of the surrounding air	A. Remains constant B. Decreases C. Increases D. Fluctuates rapidly
19	If an endothermic reaction is allowed to take place very rapidly in air, the temperature of the surrounding air	A. Remains constant B. Decreases C. Increases D. Fluctuates rapidly
20	The study of heat changes accompanying a chemical reaction is known as	A. Thermochemistry B. Biochemistry C. Physical chemistry D. Analytical chemistry
21	The net heat in a chemical reaction is same, whether it is brought about in two or more different ways in one or several steps. It is known as	A. Henry's law B. Joule's princile C. Hess's law D. law of conservation of energy
22	Which of the following statements is contrary to the first law of themodynamics	A. Energy can neither be created nor destroyed B. One form of energy can be transferred into an equivalent amount of other kinds of energy C. In an adiabatic process, the work done is independent of its path D. Continuous production of mechanical work without supplying and equivalent amount of heat is possible
23	The change in heat energy of a chemical reaction at constant temperature and pressure is called	A. entahlpy change B. heat of sublimation C. bond energy D. internal energy change
24	Calorie is equivalent to	A. 0.4184 J B. 41.84 J C. 4.184 J D. 418.4 J
25	If an endothermic reaction is allowed to take place very rapidly in the air, the temperature of the surrounding air	A. remains constant B. increases C. decreases D. remain unchanged