

ECAT Chemistry Chapter 4 Liquids & Solids Online Test

Sr	Questions	Answers Choice
1	Trend of boiling point of halogens from fluorine to Iodine is that it	A. Decreases B. Is negligible C. Increases D. Remains constant
2	Heat of vapourization for liquids with strong dipole-dipole forces will have	A. Negligible Values B. Reasonably high values C. Very high values D. very low values
3	The boiling point of NH_3 is maximum among the hydrides of group V elements due to	A. Enhanced electronegative character of Nitrogen B. Pyramidal structure of NH_3 C. Very small size of Nitrogen D. Enhanced electropositive character of Nitrogen
4	Which of the following liquids has low vapour pressure at 25°C	A. Diethyl ether B. Acetone C. Water D. Ethyl alcohol
5	Escape of high energy molecules from the surface of a liquid is called	A. Sublimation B. Distillation C. Condensation D. Evaporation
6	A liquid on evaporation causes	A. Heating effect B. Cooling effect C. Suffocation D. All of above
7	At sea level and at 100°C the vapour pressure of water in an open system is	A. 1000 mm Hg B. 760 mm Hg C. 730 mm Hg D. 670 mm Hg
8	Rate of evaporation and rate of condensation at equilibrium	A. Become very low B. Become very high C. Become equal D. Can never be equal
9	Evaporation of water is possible at	A. Above 100°C B. 0°C C. 100°C D. At all temperature
10	H_2S is a gas which H_2O is liquid at room temperature. it is due to	A. Less intermolecular forces in water B. Covalent bond in H-O in water molecule C. Hydrogen bonding in water molecules D. Ionic characters in water molecules
11	Hydrogen bonding is present between the molecules of	A. NH_3 B. H_2O C. HF D. All of above
12	Debye forces are present in one of the following pairs	A. Na^+ ion and water B. Argon and water C. Argon and Na^+ ion D. Ne and Water
13	The strongest forces are	A. Debye forces B. London dispersion forces C. Dipole-dipole attraction D. Hydrogen
14	Polypeptide chains are coiled about one another into a spiral by	A. Ionic bonds B. Covalent bonds C. Van der Waal's forces D. Hydrogen bonds

D. Hydrogen bonds

15	HF has exceptionally low acidic strengths due to	A. Smaller size of fluorine B. Strong polar bond between H and F C. Electronegativity of fluorine D. Strong hydrogen bonding
16	Hydrocarbon molecules with large chain lengths experience	A. Weaker attractive forces B. Stronger attractive forces C. Repulsive forces D. No attractive forces
17	In a group on going downward, polarizability generally	A. Decreases B. Increases C. Remains constant D. Negligible
18	London dispersion forces are also called	A. Hydrogen bonding B. Debye forces C. Van de Waal's forces D. Instantaneous dipole-induced dipole forces
19	The intermolecular forces in liquids are	A. Negligible B. Very weak C. Very strong D. Reasonably strong
20	The gases can be converted into liquids by	A. increasing the pressure only B. Lowering temperature and increasing pressure C. Increasing pressure and bringing temperature below critical point D. Lowering temperature only
21	Which one of the following is the weakest intermolecular force	A. Dipole induced dipole forces B. Ionic dipole forces C. Electrostatics forces between ions D. Dipole-dipole forces
22	The density of water decreases, when it is freezed at 0°C because of	A. Change of bond lengths B. Change of bond angles C. Cubic structure of ice D. Empty spaces present in the structure of Ice
23	The only forces are London dispersion forces among the	A. Atoms of He in gaseous state at high temperature B. Molecules of water in liquid state C. Molecules of solid $I_{2(s)}$ D. Molecular of hydrochloric acid gas
24	Force of attraction between atoms of He is	A. London dispersion forces B. Hydrogen bondign C. Coordinate covalent bond D. Covalent bond
25	Chloroform and acetone are soluble in each other due to	A. Instantaneous dipole interactions B. Dipole-dipole interactions C. Intermolecular hydrogen bonding D. All of above
26	Diamond is a bad conductor because	A. It has a tight structure B. It has a high density C. there is no free electron present in the crystal of diamond of conduct electicity D. Is transparent to light
27	Which of the following is a pseudo solid	A. CaF_2 B. Glass C. NaCl D. All
28	The molecules of CO_2 in dry ice form the	A. Ionic crystalls B. Covalent crystals C. Molecular crystals D. Any type of crystal
29	Amorphous solids	A. Have sharp melting point B. Undergo clean cleavage when cut with knife C. Have perfect arrangement of atoms D. Can possesses small regions of orderly arrangement of atoms
		A. Low melting points

- A. Low melting points
 - B. Good conductivity in solid state
 - C. High vapour pressure
 - D. Solubility in polar solvents
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