

ECAT Chemistry Chapter 4 Liquids & Solids Online Test

Sr	Questions	Answers Choice
1	The boiling point of NH_3 is maximum among the hydrides of group V elements due to:	A. Enhanced electronegative character of Nitrogen B. Pyramidal structure of NH_3 C. Very small size of Nitrogen. D. Enhanced electropositive character of Nitrogen
2	Which of the following liquids has low vapour pressure at 25°C :	A. Diethyl ether B. Acetone. C. Water. D. Ethyl alcohol.
3	Which of the following liquids has low vapor pressure at 25°C :	A. Diethyl ether B. Acetone. C. Water. D. Ethyl alcohol.
4	Escape of high energy molecules from the surface of liquid is called:	A. Sublimation. B. Distillation. C. Condensation. D. Evaporation.
5	A liquid on evaporation causes:	A. Heating effect. B. Cooling effect. C. Suffocation . D. All of above
6	At sea level and at 100°C the vapor pressure of water in an open system is:	A. 1000 mm Hg B. 760 mm Hg C. 730 mm Hg D. 670 mm Hg
7	Rate of evaporation and rate of condensation at equilibrium:	A. Become very low B. Become very high C. Become equal D. Can never be equal.
8	Evaporation of water is possible at:	A. Above 100°C B. Below 100°C C. 100°C D. At all temperature
9	H_2S is a gas while H_2O is a liquid at room temperature. It is due to:	A. Less inter-molecular forces in water. B. Covalent bonding H-O in water molecule. C. Hydrogen bonding in water molecules D. Ionic character in water molecules
10	Hydrogen bonding is present in one of the following pairs:	A. NH_3 B. H_2 C. H_2O D. CH_4

		<p>C. H^+</p> <p>D. All of above</p>
11	Debye forces are present in on of the following pairs:	<p>A. Na^+ion and water</p> <p>B. Argon and water</p> <p>C. Argon and Na^+ion</p> <p>D. Ne and water</p>
12	The strongest forces are:	<p>A. Debye froces</p> <p>B. London dispersion</p> <p>C. Dipole-dipole attraction</p> <p>D. Hydrogen bonding</p>
13	Polypeptide chains are coiled about one another into a spiral by:	<p>A. Ionic bonds</p> <p>B. Covalent bonds</p> <p>C. Van der Waal's forces</p> <p>D. Hydrogen bonds</p>
14	HF has exceptionally low acidic strength due to:	<p>A. Smaller size of fluorine.</p> <p>B. Stronger polar bond between H and F.</p> <p>C. Electronegativity of fluorine</p> <p>D. Strong hydrogen bonding</p>
15	Hydrocarbon molecule with large chain lengths experience:	<p>A. Weaker attractive forces</p> <p>B. Stronger attractive forces</p> <p>C. Repulsive forces</p> <p>D. No attractive forces</p>
16	In a group on going downward, polarizability generally:	<p>A. Decreases</p> <p>B. Increases</p> <p>C. Remain constant</p> <p>D. Negligible</p>
17	London dispersion forces are also called:	<p>A. Hydrogen bonding.</p> <p>B. Debye forces</p> <p>C. Van der Waal's forces</p> <p>D. Instantaneous dipole-induced dipole forces.</p>
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19	The inter-molecular forces in liquids are:	<p>A. Negligible</p> <p>B. Very weak</p> <p>C. Very strong</p> <p>D. Reasonably strong</p>
20	The gases can be converted into liquids by:	<p>A. Increasing the pressure only.</p> <p>B. Lowering temperature and increasing pressure</p> <p>C. Increasing pressure and bringing temperature below critical points</p> <p>D. Lowering temperature only</p>
21	Which one of the following is weakest inter molecular force?	<p>A. Dipole induced dipole forces</p> <p>B. Ionic dipole forces</p> <p>C. Electrostatic forces b/w ions</p> <p>D. Dipole dipole forces</p>
22	The density of water decreases, When it is freezed at 0°C	<p>A. Change of bond length</p> <p>B. Change of bond angles</p> <p>C. Cubic structure of ice</p> <p>D. Empty spaces present in the structure of ice</p>
23]Which of the following is a pseudo solid?	<p>A. Atoms of He is gaseous stat at high temperate.</p> <p>B. Molecules of water in liquid state.</p> <p>C. Molecules of solid $\text{I}_{2(s)}$</p> <p>D. Molecular of hydrochloric acid gas</p>
24	Force of attraction b/w atoms of He is :	<p>A. London dispersion forces</p> <p>B. Hydrogen bonding</p> <p>C. Coordinate covalent bond</p> <p>D. Covalent bond</p>
25	Chloroform and acetone are soluble in each other due to:	<p>A. Instantaneous dipole interactions.</p> <p>B. Dipole-dipole interactions.</p> <p>C. Inter molecular hydrogen bonding.</p> <p>D. All of above</p>
26	Diamond is a bad conductor because:	<p>A. It has tight structure</p> <p>B. It has a high density</p> <p>C. There is no free electron present in the crystal of diamond of conduct</p>

27	Which of the following is a pseudo solid?	<p>A. CaF_2</p> <p>B. Glass</p> <p>C. NaCl</p> <p>D. All</p>
28	The molecules of CO_2 in dry ice form the :	<p>A. Ionic crystals</p> <p>B. Covalent crystals</p> <p>C. Molecular crystals</p> <p>D. Any type of crystals</p>
29	Amorphous solids:	<p>A. Have a sharp melting point</p> <p>B. Undergo clean cleavage when cut with knife</p> <p>C. Have a perfect arrangement of atoms</p> <p>D. Can possess small regions of orderly arrangement of atoms</p>
30	Ionic solids are characterized by :	<p>A. Low melting points</p> <p>B. Good conductivity in solid state</p> <p>C. High vapor pressure</p> <p>D. Solubility in polar solvents</p>