

## ECAT Chemistry Chapter 1 Basic Concepts Online Test

Sr	Questions	Answers Choice
1	Benzene is stable to:	A. Oxidation B. Nitration C. KMnO <sub>4</sub> D. SULPHONATION
2	Formation of a cation is:	A. Exothermic process B. Non-endothermic process C. Endothermic process D. None of above
3	An ion bearing positive charge is called:	A. Cation B. Positron C. Anion D. None of above
4	A species having positive or negative charge is called:	A. Electron B. Ion C. Proton D. Atom
5	Which statement about molecule is incorrect ?	A. Molecules of a substance are similar B. Hemoglobin is a homo atomic molecules C. Oxygen molecule is a macro molecule D. It exist independently
6	Which statement about an atom is true ?	A. The number of neutrons is not equal to number of electrons B. Mass number is less than atomic number C. All the elements have only one mass number D. Mass number can be equal to atomic number
7	In molecules kinetic and potential energies are:	A. Definite B. Moderate C. Indefinite D. None of above
8	Molecules of High molecular weight usually greater than 10,000 are called:	A. Macro molecules B. Mega molecules C. Poly molecules D. Gega molecules
9	Hemoglobin contains nearly:	A. 10,000 atoms B. 100 atoms C. 1000 atoms D. 1 atom
10	Hemoglobin is 68000 times heavier than:	A. Oxygen atom B. Nitrogen atom C. Carbon atom D. Hydrogen atom
		A. <p class="MsoNormal"><span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Diatomic molecules</span><o:p></o:p></p> B. <p class="MsoNormal"><span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Hemoglobin</span><o:p></o:p></p>

11 NH<sub>3</sub>, HCl, H<sub>2</sub>O, HL are:

- A. <p class="MsoNormal"><span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Poly-atomic molecules</span><o:p></o:p></p>
- C. <p class="MsoNormal"><span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Mono-atomic molecules</span><b><o:p></o:p></b></p>
- D. <p class="MsoNormal"><span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Heter</span>o atomic molecules<o:p></o:p></p>

12 He Ar and Ne are:

- A. <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Mono-atomic molecules</span>
- B. <p class="MsoNormal"><span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Heter</span>o atomic molecules<o:p></o:p></p>
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- D. <p class="MsoNormal"><span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Diatomc molecules</span><o:p></o:p></p>

13 CL<sub>2</sub>, N<sub>2</sub> and O<sub>2</sub> are:

- A. <p class="MsoNormal"><span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Diatomic molecules</span><o:p></o:p></p>
- B. <p class="MsoNormal"><span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Heter</span>o atomic molecules<o:p></o:p></p>
- C. <p class="MsoNormal"><span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">

14 C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>and C<sub>12</sub>H<sub>22</sub>O<sub>11</sub> are:

- A. Mono-atomic molecules  
B. Diatomic molecules  
**C. Poly-atomic molecules**  
D. <p class="MsoNormal"><span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Poly-atomic molecules</span><o:p></o:p></p>  
D. <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Mono-atomic molecules</span>

15 The number of subatomic particles in atoms sidcovered is more than:

- A. 110  
**B. 100**  
C. 125  
D. 90

16 Atoms can be evident by use of electron microscope, field ionization microscope and:

- A. x-rays**  
B. Video camera<div><br></div>  
C. Telescope  
D. Compound microscope

17 Determination of atomic masses and invention of system of writing symbols was made by:

- A. J. Berzelius**  
B. Democritus  
C. Dalton  
D. None of above

18 First atomic theory was put forward by an English school teacher:

- A. Maxwell  
B. Newton  
C. Sanger  
**D. John Dalton**

19 Metal tend to lose electrons, becoming:

- A. <p class="MsoNormal">Metals</o:p></o:p></p>  
B. <p class="MsoNormal">Positively charged</o:p></o:p></p>  
**C. <p class="MsoNormal">Negatively charged</o:p></o:p></p>**  
D. <p class="MsoListParagraph" style="text-indent:-.25in;mso-list:l0 level1 lfo1"><span style="font-variant-numeric: normal; font-variant-east-asian: normal; font-stretch: normal; line-height: normal;"> (a)</span><span style="font-variant-numeric: normal; font-variant-east-asian: normal; font-stretch: normal; font-size: 7pt; line-height: normal; font-family: &quot;Times New Roman&quot;"></span><!--[endif]-->And (c)<o:p></o:p></p>

20 Atoms and molecules can either gain or lose electrons, forming charge particles called:

- A. <p class="MsoNormal">Positrons</o:p></o:p></p>**  
B. <p class="MsoNormal">Photons</o:p></o:p></p>  
**C. <p class="MsoNormal">Ions</o:p></o:p></p>**  
D. <p class="MsoNormal">Electrons</o:p></o:p></p>

- A. <p class="MsoNormal">Protons</o:p></o:p></p>**  
B. <p class="MsoNormal">Atoms</o:p></o:p></p>

- 21 Covalent compound s mostly exist in the form of:  
C. <span style="font-size:11.0pt;line-height:107%; font-family:&quot;Calibri&quot;,&quot;sans-serif&quot;;mso-ascii-theme-font:minor-latin;mso-fareast-theme-family: Calibri;mso-fareast-theme-font:minor-latin;mso-hansi-theme-font:minor-latin; mso-bidi-font-family:&quot;Times New Roman&quot;;mso-bidi-theme-font:minor-bidi; mso-ansi-language:EN-US;mso-fareast-language:EN-US;mso-bidi-language:AR-SA">Neutrons</span>  
D. <p class="MsoNormal">Molecules<o:p></o:p></p>
- 
- 22 The diameter of atoms is of the order:  
A.  $2 \times 10^{-5}$ m  
B. <p class="MsoNormal"> $2 \times 10^{10}$  m<sup>-5</sup><o:p></o:p></p>  
C.  $2 \times 10^{-2}$ m  
D.  $2 \times 10^{-3}$ m
- 
- 23 CO<sup>+</sup> is an example of  
A. Stable molecule  
B. Anionic molecule ion  
C. Cationic molecular ion  
D. Free radical
- 
- 24 When an electron is added to a uni positive ion we get:  
A. Cation  
B. Molecule  
C. Neutral atom  
D. Anion
- 
- 25 The number of atoms present in molecule determines its:  
A. Molecularity  
B. Atomicity  
C. Basicity  
D. Acidity
- 
- 26 Matter is defined as any thing which occupies space and:  
A. Molecules  
B. Mass  
C. Compound  
D. Molecules
- 
- 27 Smallest particle of an element which may or may not have independent existence is known as:  
A. A molecule  
B. An ion  
C. An atom  
D. An electron
- 
- 28 The branch of science dealing with structure, composition and changes in matter and laws and principles which govern these changes is called as  
A. chemistry  
B. Geology  
C. Physics  
D. Mechanics
- 
- 29 Isotopes differ in  
A. properties which depend upon mass  
B. arrangement of electrons in orbitals  
C. chemical properties  
D. the extent to which they may be affected in electric fields
- 
- 30 The number of isotopes of gold is  
A. 3  
B. 1  
C. 2  
D. 4