ECAT Chemistry Chapter 12 Periodic Classification of Elements and Periodicity Online Test

Sr	Questions	Answers Choice
1	Among the elements given below, the one with highest electropositivity is	A. Cu B. Cs C. Cr D. Ba
2	The correct arrangement of increasing order of atomic radius among Na, K, Mg, Rb is	A. Mg < K < Na < Rb B. Mg < Na < K < Rb C. Mg < Na < Rb < K D. Na < K < Rb < Mg
3	Alkali metals in each period have	A. Smallest size B. Lowest IE C. Highest IE D. Highest electronegativity
4	The element with atomic number 55 belongs to which block of the periodic table	A. s-block B. p-block C. d-block D. f-block
5	The attraction that an atom exerts on a pair of electrons that are being shared with another atom for forming covalent bond is referred to as its	A. Electron affinity B. Electronegativity C. Ionisation energy D. Valency
6	Among O, C, F, Cl, Br, the correct order of increasing radii is	A. F O C Cl Br B. F C O Cl Br C. F Cl Br O C D. C O F Cl Br
7	Among the following elements which one has the highest value of first ionization potential?	A. Oxygen B. Argon C. Barium D. Cesium
8	The element with atomic number 26 will be found in group	A. 2 B. 8 C. 6 D. 10
9	Which of the following pair of atomic numbers represents s-block elements?	A. 7, 15 B. 6, 12 C. 9, 17 D. 3, 20
10	The valence shell electronic structure of an element is ns^2np^5. The element will along to the group of	A. Alkali metals B. Inert metals C. Noble gases D. Halogen
11	The valency of noble gases, in general, is	A. Zero B. One C. Three D. Two
12	Which among the following elements have lowest value of IE_1?	A. Pb B. Sn C. Si D. C
13	Of the given alkali metals, the one with smallest size is	A. Rb B. Cs C. K D. Na
14	Which of the following elements is/are not liquid at 30°C?	A. Ga B. Hg C. Ge D. Cs
15	Which of the following metal requires radiation of highest frequency to cause emission of electrons?	A. Na B. Mg C. K D. Ca

16	Which of the following does not reflect the periodicity of elements?	A. Bonding behaviour B. Electronegativity C. Ionisation potential D. Neutral/proton ratio
17	Which of the following has greatest tendency to lose electron?	A. F B. Fr C. S D. Be
18	Which of the following elements is most electronegative?	A. Oxygen B. Chlorine C. Nitrogen D. Fluorine
19	The alkali metal which is liquid at 15°C is	A. K B. Cs C. Na D. None
20	Which of the following pairs are chemically dissimilar?	A. Na and K B. Ba and Sr C. Zr and Hf D. Ca and Zn
21	Variable valency is generally exhibited by	A. Normal elements B. Transition elements C. Metallic elements D. None of these
22	Which of the following statements is most appropriate about effective nuclear charge? It depends upon	A. The shielding constant B. The atomic number C. The charge on the nucleus D. Both the nuclear charge and the shielding constant
23	Which of the following represents elements in order of increasing atomic size?	A. I, Br, Cl B. Na, Mg, C C. C, N, O D. Li, Na, K
24	The chloride of element Q is hydrolysed by water to form an acidic solution and its oxide reacts with acid to form a salt. What cold be the element Q	A. Magnesium B. Aluminium C. Silicon D. Phosphorus
25	Which statement explains the observation that magnesium hydroxide dissolve in aqueous ammonium chloride, but not in aqueous sodium chloride	A. The ionic radius of the NH₄⁺ion is similar to that of Mg²⁺but not that of Na⁺ B. NH₄Cl dissociates less fully than NaCl C. The ions Na⁺and Mg²⁺are isoelectronic (have the same number of electrons) D. The ion NH⁺₄acts as an acid
26	An element of the third period (Na to S) is heated in chlorine. The product is purified and then added to water. The resulting solution is found to be neutral. What is the element	A. Sodium B. Aluminium C. Silicon D. Phosphorus
27	Which of the following oxides is unlikely to dissolve in aqueous hydroxide	A. Al₂O₃ B. MgO C. SO₂ D. SlO₂
28	Which species represented by the following formula has the largest radius	A. P³⁻ B. Cl⁻ C. A_r D. K⁺
29	$NaBH_4$and $LialH_4$are	A. Ionic hydrides B. Covalent hydrides C. Interposal hydrides D. Complex hydrides
30	Ionization potential increases in moving from left to right in a period	A. Because nuclear charge increase B. Because atomic size decrease C. Both (a) and (b) D. Because atomic size increases