ECAT Chemistry Chapter 12 Periodic Classification of Elements and Periodicity Online Test

Sr	Questions	Answers Choice
1	The melting point is lowest for	A. Be B. Mg C. Ca D. Sr
2	The number of elements in the 4th periods of periodic table is	A. 8 B. 10 C. 18 D. 32
3	The oxides of which of the following elements will be acidic in character	A. Mg B. Rb C. Li D. Cl
4	Which of the following is an inert gas?	A. H₂ B. O₂ C. N₂ D. Argon
5	Which of the following oxides is amopheric in character?	A. CaO B. CO₂ C. SiO₂ D. SnO₂
6	Which of the following ion has the highest value of ionic radius?	A. Li⁺ B. F⁻ C. O²⁻ D. B³⁺
7	The elements with atomic numbers 9, 17, 35, 53, 85 an all	A. Noble gases B. Halogens C. Heavy metals D. Light metals
8	How does the ionization energy of 1st group elements vary?	A. Increases down the group B. Decreases down the group C. Remains unchanged D. Variation is not regular
9	Number of elements present in 5th period is	A. 8 B. 18 C. 32 D. 24
10	Two elements whose electronegativities are 1.2 and 3.0, the formed between them would be	A. Ionic B. Covalent C. Coordinate D. Metallic
11	Electron affinity depends on	A. Atomic size B. Nuclear charge C. Atomic number D. Atomic size and nuclear charge both
12	Which is true about the electronegativity order of the following?	A. P > Si B. C > N C. Br > Cl D. Sr > Ca
13	Eka-aluminium and Eka-silicon are known as	A. Gallium and Germanium B. Aluminium and silicon C. Iron and sulphur D. Proton and silicon
14	Which among the following species has the highest ionization energy?	A. Ne B. F C. Li D. B
15	Which of the following does not exhibit the periodicity in properties of the elements?	A. Ionisation energy B. N/P ratio C. Electronegativity

		D. Atomic radius
16	Which of the following sets of atomic numbers belong to that of the alkali metals?	A. 1,12,30,4,62 B. 37,19,3,55 C. 9,17,35,53 D. 12,20,56,88
17	Which of the following isoelectronic ions has the lowest ionization energy?	A. K⁺ B. Ca²⁺ C. Cl⁻ D. S²⁻
18	The atomic radius increases as we move down a group because	A. Effective nuclear charge increases B. Atomic mass increases C. Additive electrons are accommodated in new electron level D. Atomic number increases
19	Which of the following statement about fluorine is not correct?	A. Electron affinity of chlorine is greater than that of fluorine B. Bond energy of fluorine is less than that of chlorine C. Fluorine cannot be prepared by electrolysis of fused metal fluorides D. Fluorine does not form oxoacid
20	The element with highest electron affinity among the halogen is	A. F B. Cl C. Br D. I
21	The ionization potential is lowest for the	A. Halogens B. Inert gases C. Alkaline earth metals D. Alkali metals
22	In the modern long form of the periodic table elements are arranged in the increasing order of	A. Atomic mass B. Atomic number C. Mass number D. Isotopic number
23	Variable valency is characteristic of	A. Halogen B. Transition elements C. Alkali metals D. Noble gas
24	The correct order of electron affinity among the following is	A. F > Cl > Br B. Br > Cl > F C. Cl > F > Br D. F > Br > Cl
25	Which of the following is most electronegative?	A. Carbon B. Silicon C. Lead D. Tin
26	Which of the following element has the maximum electron affinity?	A. F B. S C. I D. Cl
27	Which of the following has highest first ionization potential?	A. Carbon B. Oxygen C. Nitrogen D. Boron
28	Which of the following species has the highest ionization potential?	A. Ne B. Al⁺ C. Mg⁺ D. Li⁺
29	Gradual addition of electronic shells in the nobel gases causes a decrease in their	A. Ionization energy B. Atomic radius C. Boiling point D. Density
30	In the periodic table, the element with atomic number 16 will be placed in the group	A. Fourteen B. Sixteen C. Thirteen D. Fifteen