

ECAT Chemistry Chapter 12 Periodic Classification of Elements and Periodicity Online Test

Sr	Questions	Answers Choice
1	Which element should have the greatest value for electronegativity when combined with hydrogen	A. Na B. Si C. S D. Cl
2	Which of the following elements have the largest radius	A. F B. Cl C. Br D. I
3	Which of the following elements should be the least metallic in character	A. Rb B. In C. Te D. I
4	Gradation in properties in the periods of periodic tables are due to change in	A. Atomic weight B. The number of electrons C. Number of protons D. Electronic configuration
5	The most distinctive character among the elements is their division into	A. Metals and non-metals B. Solids, liquids and gases C. Atoms and molecules D. Active and inactive elements
6	Which of the following discoveries resulted in a version of the Mendeleeffs periodic law	A. The nucleus of atom by Rutherford B. The elements polonium and radium by the Curies C. Atomic numbers by Moseley D. x-rays by Roentgen
7	The hydration energy is the heat evolved when one mole of gaseous ion is dissolved in water. The hydration energy of an ion	A. Increases with increase of charge to mass ratio B. Decreases with increase of charge to mass ratio C. Depends on sign of charge +ve or -ve D. Depends upon the solvent
8	Although hydrogen resemble with the elements of group IA, IVA and VII but it is usually placed in	A. Group IA B. Group IV A C. Group VII D. Group VIII
9	Which of the following is a member of -block	A. Zn B. Al C. B D. Br
10	The fourteen elements following actinium are known as	A. Lanthanones B. Lanthanides C. Rare earths D. Actinides
11	The fourth period contains elements	A. 8 B. 16 C. 18 D. 32
12	What is the nature of Al_2O_3	A. Acidic B. Basic C. Amphoteric D. Neutral
13	What is the nature of SO_2	A. Basic B. Strongly acidic C. Weakly acidic D. Amphoteric
14	Na_2O is	A. Acidic B. Basic C. Neutral D. Amphoteric

15	A pair of elements in the same family in the periodic classification is	A. Cl and C B. Ca and Al C. N and Ne D. Na and K
16	The elements of group IA are called	A. Chalcogens B. Halogens C. Alkali metals D. Alkaline earth metals
17	The oxides of electronegative elements are	A. Basic B. neutral C. Acidic D. Amphoteric
18	The coinage metals are	A. Ni, Pd, Pt B. Cu, Ag, Au C. Zn, Al, Pb D. Fe, Si, Sn
19	In a period, melting points of elements	A. Increases B. Decreases C. Remain constant D. First increases then decreases
20	Ionization energy depends upon	A. Nuclear charge B. Atomic size C. Shielding effect D. All of the above
21	In a group, the ionization energy	A. Increase B. Decreases C. Remain constant D. First increases then decreases
22	The amount of energy required to remove an electron from an atom of an element in the gaseous state is called	A. Electron affinity B. Electronegativity C. Ionization energy D. None of these
23	Which is the transition element among the following	A. B B. Al C. Cu D. Cs
24	In sixth period 14 of its transition elements are called	A. Lanthanides B. Actinides C. Radioactive elements D. None
25	The correct order of 2nd I.P. of C, N, O and F is	A. O > F > N > C B. O > N > F > C C. C > N > O > F D. F > O > N > C
26	Which is the longest period of the periodic table	A. 5th B. 7th C. 6th D. 2nd
27	Which orbital is in the process of completion in case of transition elements	A. p-orbital B. f-orbital C. d-orbital D. s-orbital
28	Number of elements in the first period of the periodic table are	A. Two B. Four C. One D. Eight
29	The energy absorbed when an electron is added to a gaseous atom to form a gaseous ion is called	A. Electron affinity B. Ionization energy C. Both of these D. None of these
30	The positive ion is always smaller than the neutral atom while the negative ion is always bigger than the neutral atom. The atomic and ionic radii of Na, F, Na ⁺ , F ⁻ are in pm	A. Na F Na ⁺ F ⁻ 157 72 95 136 B. Na F Na ⁺ F ⁻ 157 95 172 136 C. Na F Na ⁺ F ⁻ 72 95 136 157 D. Na F Na ⁺ F ⁻ 157 136 95 72