ECAT Chemistry Chapter 11 Reaction Kinetics Online Test

Sr	Questions	Answers Choice
1	A white precipitate ofsilver chlorideimmediately formed on additionof :	A. Silver nitrate solution to sodium chloride solution. B. Silver chloride solution to sodium nitrate solution. C. Silver nitrate solution to potassium chloride solution D. Silver nitrate solution to hydrogen chloride solution.
2	A white precipitate ofsilver chlorideimmediately formed on additionof :	A. Silver nitrate solution to sodium chloride solution. B. Silver chloride solution to sodium nitrate solution. C. Silver nitrate solution to potassium chloride solution D. Silver nitrate solution to hydrogen chloride solution.
3	Question Image	A. 2 B. 3 C. 4 D. 9
4	If the rate of reaction is independent of the concentration of the reactant, the reaction is of	A. Zero order B. First order C. Second order D. Third order
5	The addition of a catalyst to a reaction changes the	A. Enthalpy B. Entropy C. Nature of reactants D. Energy of activation
6	Which of the following factors does not influenced the rate of reaction	A. Concentration of the reaction B. Nature of the reactants C. Molecularity of the reaction D. Temperature
7	The experimental relationship between a reaction rate and the concentration of reactants is known as	A. Order B. Molecularity C. Rate constant D. Rate law
8	The number of atoms or molecules whose concentrations determine the rate of the reaction is called	A. Molecularity B. Order C. Rate of reaction D. Rate constant
9	_____ are called biocatalysts	A. Organic acids B. Organic bases C. Enzymes D. All
10	Complex protein molecules which catalyses the organic reactions in the living cells are called	A. Living organisms B. Enzymes C. Viruses D. Bacteria
11	Question Image	A. Diastase B. Lipase C. Inverters D. Zymase
12	When copper is allowed to react with HNO_3, the reaction is slow in the beginning, finally becomes very fast. It is due to the formation of an auto catalyst which is	A. Cu(NO₃)₂ B. CuO C. O₂ D. HNO₂
13	In the reaction of oxalic acid with $KMnO_4$and $H_2SO_4$is slow at the beginning but after sometimes the reaction becomes faster due to	A. Formation of MnSO₄which acts as 'Auto catalyst B. Formation of CO₂which acts as 'Auto catalyst C. Formation of

		K₂SO₄which acts as 'Auto catalyst D. Evolution of O₂gas which acts as 'Auto catalyst
14	Hydrogenation of vegetable oils is accelerated by Ni catalyst. The catalytic activity of Bi is increased by a promoter of activator which is	A. Na and K B. Na and Hg C. Hg and Zn D. Cu and Te
15	In the manufacture of $NH_3$by Haber's process catalyst used is iron its catalytic efficiency is poisoned by	A. Presence of Al₂O₃ B. Presence of Cr₂O₃ C. MnO₂ D. CO present with H₂gas
16	The catalytic activity of Pt is much higher when	A. It is mixed with asbestos B. It is mixed with Pd C. It is mixed with arsenic D. In is made colloidal platinum
17	The effective activity of a metal catalyst is increased if it is in	A. Solid form B. Liquid state C. Gaseous state D. Finely divided form
18	Question Image	A. Homogeneous B. Heterogeneous C. Isogeneous D. None
19	Platinum is poisoned by	A. Arsenic B. Silver C. Argon D. Zinc
20	which one of the following is a heteroheneous catalysis	
21	Homogenous catalysis is that in which catalyst and reactants are in same phase. Which one of the following reaction is a homogenous catalysis	
22	A catalyst is a substance which increase the rate of a chemical reaction, but remains unchanged at the end of reaction, nut remains unchanged at the end of reaction, because	A. It increases the temperature B. It increase the surface area C. It increases the rate constant D. It decrease the energy energy of activation
23	The value of activation energy Ea of a reaction can be determined from the value of slope of the straight line obtained by plotting a graph between 1/T and log k. the value of Ea is equal to	A. Slope B. 1/Slope C. Slope x R D. Slope x 2.303 R
24	Which statement about Arrhenius equation is incorrect	A. Factor 'A' called Arrhenius constant depends upon collision frequency of reactants B. Rate of reaction increase by increasing temperature C. Rate constant k is increased D. Activation energy Ea is decreased by rise in temperature
25	The rate constant k of a reaction activation energy Ea and temperature are related by Arrhenius in the form of an equation which is	
26	With increases in temperature of 10 K of the reacting gases the rate of reaction is doubled because	A. Increase in number of collisions B. Number of molecules having energy more than Ea is doubled C. Increase in order of reaction D. Increase in surface area
27	The factor which effect the rate of reaction	A. Nature of reactants B. Surface area C. Light D. All of the above
28	Which of the following is not affected by light	
29	Which one of the following reaction rate is effected by the light	
30	The actual number of atoms or molecules taking part in rate determining step is	A. Rate of reaction B. Velocity of reaction C. Order of reaction D. Molecularly