

ECAT Chemistry Chapter 11 Reaction Kinetics Online Test

Sr	Questions	Answers Choice
1	The unit of rate constant is the same as that of the rate of reaction in :	A. First order reaction. B. Second order reaction. C. Zero order reaction. D. Third order reaction.
2	With increases of 10°C temperature the rate of reaction doubles. This increase in rate of reaction is due to :	A. Decrease in activation energy or reaction. B. Decrease in number of collisions between reactant molecules. C. Increase in activation energy of reactants. D. Increase in number of effective collisions
3	The rate of reaction :	A. Increase as the reaction proceeds. B. Decreases as the reaction proceeds. C. Remains the same as the reaction proceed. D. May decrease or increase as the reaction proceeds.
4	If the rate equation of a reaction $2A + B \rightarrow \text{Products}$ is , Rate = $K[A][B]$, and A is present in large excess, then order of reaction is :	A. 1 B. 2 C. 3 D. Above
5	In zero order reaction, the rate is independent of :	A. Temperature of reaction B. Concentration of reactants C. Concentration of products D. None of these.
6	Which technique is used to determine the absorption of radiations?	A. <p>class="MsoNormal">Dilatometer method</p></p> B. <p>class="MsoNormal">Optical rotation method</p></p> C. <p>class="MsoNormal">Spectrometry</p></p> D. <p>class="MsoNormal">Refractometric method</p></p></p></p></p></p>
7	The rate of reaction :	A. Decreases as the reaction proceeds B. Increases as the reaction proceeds C. May decrease or increase reaction proceeds D. Remains same as the reaction proceeds
8	Dilatometer method is useful for the reaction that involve :	A. Small volume changes in solutions B. Change in refractive indices C. Where reactants absorb U.V, visible or infrared radiation
9	The rate of reaction b/w two specific time intervals is called :	A. Instantaneous rate of reaction. B. Average rate of reaction. C. Rate of a reaction. D. Minimum rate of a reaction.
10	If the rate of decay of radioactive isotope decreases from 200 cpm to 25 cpm after 24 hours, what is its half life :	A. 8 hours B. 6 hours C. 4 hours D. 3 hours
11	Half life period of a first order reaction is independent of:	A. Presence of catalyst. B. Conditions of temperature C. Initial concentration of the compound D. All of above
12	The unit of the rate constant is the same as that of rate of reaction in :	A. Third order reaction B. Second order reaction C. First order reaction D. Zero order reaction

13	The rate equation for a reaction is $\text{Rate} = k[A]$. what are unit of K ?	<p>A. $\text{Mole-dm}^{-3} \text{S}^{-1}$</p> <p>B. $\text{Mole-dm}^{-3} \text{S}^{-1}$</p> <p>C. $\text{Mole-dm}^{-3} \text{S}^{-1}$</p> <p>D. S^{-1}</p>
14	A zero order reaction is one in which :	<p>A. Rate is not affected by changing concentration of reactants.</p> <p>B. concentration of reactants do not change with the passage of time.</p> <p>C. Reactants do not react.</p> <p>D. One reactants in large excess.</p>
15	A pseudo uni-molecular reaction has order of reaction :	<p>A. 3</p> <p>B. 2</p> <p>C. 1</p> <p>D. 0</p>
16	Rate law of an equation is obtained :	<p>A. From a balance equation.</p> <p>B. Can be calculated theoretically as well as determined experimentally.</p> <p>C. It is only calculated theoretically.</p> <p>D. Experimentally.</p>
17	The example of a photo chemical reaction is photosynthesis has order of reaction :	<p>A. 1</p> <p>B. 2</p> <p>C. 0</p> <p>D. 3</p>
18	It rate law of an equation is written $\frac{dx}{dt} = k[A][B]$?	<p>A. Reaction is independent of the concentration of A and B.</p> <p>B. Product is decreasing with passage of time.</p> <p>C. Reactant is increasing with passage of time.</p> <p>D. Reactant is increasing with passage of time.</p>
19	Hydrolysis of ethyl-acetate (ester) has order of reaction :	<p>A. 3</p> <p>B. 2</p> <p>C. 1</p> <p>D. 1</p>
20	The unit rate of constan K is $\text{mole}^{-1} \text{dm}^3 \text{S}^{-1}$ for a chemical reaction, the order of reaction is :	<p>A. 3</p> <p>B. 2</p> <p>C. 1</p> <p>D. 0</p>
21	The unit of rate constant K is $\text{mole}^{-1} \text{dm}^3$ for a chemical reaction, the order of reaction is :	<p>A. Order of reaction can determined by an experiment</p> <p>B. Order of reaction can determined from a balance equation only.</p> <p>C. Order of reaction can determined increases by increasing temperate.</p> <p>D. Order of reaction must be in whole number and not in fraction.</p>
22	Which statement is true about order of reaction :	<p>A. Order of reaction can only be determined by an experiment.</p> <p>B. Order of reaction can be determined from a balance equation only.</p> <p>C. Order of reaction increase by increasing temperature.</p> <p>D. Order of reaction must be in whole number and not in fraction.</p>
23	Rate of chemical reaction depends upon :	<p>A. The number of total collisions per second.</p> <p>B. Number of molecules taking part in a chemical reaction.</p> <p>C. Number of fruitful collisions per second</p> <p>D. Number of fruitless collisions per second.</p>
24	The change in concentration of reactant or product per unit time is called :	<p>A. Rate constant.</p> <p>B. Rate of reaction.</p> <p>C. Rate equation.</p> <p>D. Rate law.</p>
25	Which of the following may affect the rate constant (k) fro a reaction :	<p>A. Change in concentration.</p> <p>B. Change in pressure.</p> <p>C. Change in pH</p>

		C. Change in pH. D. Change in temperature.
26	It is common observation that rates of chemical reactions differ :	A. Greatly. B. A little bit. C. Moderately.
27	Which of the following will affect the rate :	A. First step of reaction. B. Last step of reaction. C. Rate determining step. D. Fastest step.
28	The rate determining step is the :	A. Slowest step. B. Fastest step. C. Moderate step. D. Both (a) and (b).
29	All reactions occur in :	A. A single step. B. A series of steps C. Two steps. D. Both (a) and (b)
30	Which of the following reactions occur at moderate rate :	A. Rusting of iron B. Chemical weathering of stone work of buildings by acidic gases in atmosphere. C. Hydrolysis of an ester D. Fermentation of sugars