

## ECAT Chemistry Chapter 11 Reaction Kinetics Online Test

Sr	Questions	Answers Choice
1	The experimental relationship between a reaction rate and the concentration of reactants is called	A. Order of reaction B. Rate law C. Activated complex D. Molecularity
2	Question Image	A. $\text{Rate} = k[\text{FeCl}]^3 [\text{KI}]^2$ B. $\text{Rate} = k[\text{Fe}]^{+3} [\text{Cl}]^{-1} [\text{KI}]$ C. $\text{Rate} = k[\text{Fe}]^{+3} [\text{Cl}]^{-1} [\text{K}_2]$ D. $\text{Rate} = k[\text{KI}]^3 [\text{FeCl}]^3$
3	Question Image	A. First order B. Pseudo first order C. Second order D. Zero order
4	Question Image	A. Zero B. 1 C. 2 D. 1.5
5	The rate of reaction is denoted by	A. $\frac{dc}{dp}$ B. $\frac{dc}{ac}$ C. $\frac{dc}{dT}$ D. $\frac{dc}{dt}$
6	The sum of the exponents of the conc. terms in the rate equation is called	A. Rate of reaction B. Order of reaction C. Specific rate constant D. Average rate
7	In the hydrolysis of $\text{CH}_3\text{COOC}_2\text{H}_5$ the acid produced is	A. Inhibitor B. Catalyst C. Auto catalyst D. None of above
8	The rate of reaction between two specific time intervals is called	A. Instantaneous rate B. Average rate C. Specific rate D. Ordinary rate
9	The rate of reaction determined at a given time is called	A. Average rate B. Instantaneous rate C. Specific rate D. Overall rate
10	In the rate equation when the concentration of reactants are unity, then rate is equal to	A. Instantaneous rate B. Average rate C. Active mass of products D. Specific rate constant
11	Question Image	A. Small change in concentration of product B. Small time interval C. Co-efficient of the reactant D. Co-efficient of the product
12	The unit of rate of reaction is	A. $\text{mole dm}^{-3}$ B. $\text{mole Kg}^{-1}$ C. $\text{moles dm}^{-3} \text{sec}^{-1}$ D. $\text{grams dm}^{-3}$
13	Which statement is not correct	A. Enzymes catalyst a specific reaction B. Enzymes show catalytic activity at a specific temperature C. The catalytic activity of enzymes is stopped if optimum pH is changed D. The catalytic activity is poisoned by a co-enzymes

14	A substance which increases the rate of a reaction without being consumed during the reaction is called	A. An autocatalyst <b>B. A catalyst</b> C. A negative catalyst D. All of these
15	Factor which slows down the rate of reaction is	A. Small size of the particles of the reactant B. High temperature of reaction C. More concentration of reactant <b>D. Lowering the temperature</b>
16	Activation energy is the difference of energy between the energy of the reactant and	A. The product <b>B. The activated complex</b> C. Both a and b D. None of these
17	The chemical method used for determination of rate of reaction is	A. Spectroscopic B. Conductimetric C. Refractometric <b>D. Titration</b>
18	Question Image	A. Zero <b>B. 253 sec</b> C. 150 sec D. 500 sec
19	In the expression rate = $K[A]^a[B]^b$ K is	A. The order of reaction B. The speed of reaction <b>C. The specific rate constant</b> D. The overall order of reaction
20	The rate of reaction determined at a given time is called	A. Average rate <b>B. Instantaneous rate</b> C. Specific rate D. Overall rate
21	With the progressive of the reaction the slope of the curve between concentration of product and time	A. Gradually becomes more steep <b>B. Gradually becomes less steep</b> C. No change occurs in slope D. None of these occurs
22	The reaction rate is expressed in the units of	<b>A. <math>\text{Mol dm}^{-3}\text{s}^{-1}</math></b> B. $\text{Mol dm}^{-3}$ C. $\text{Mol dm}^{-3}\text{N}^{-1}\text{s}^{-1}$ D. $\text{dm}^{-3}\text{s}^{-1}$