ECAT Chemistry Chapter 10 Electrochemistry Online Test

Sr	Questions	Answers Choice
1	In electronic watches or electronic calculators the tiny batteries used are	A. Alkaline battery B. NICAD battery C. Fuel cell D. Silver oxide battery
2	A dry alkaline cell has porous Zn anode and MnO_2as cathode the electrolyte used is	A. Ca(OH)₂ B. NaOH C. KOH D. NH₄OH
3	Question Image	A. Fe is reduced B. Fe is oxidized C. Cl₂is oxidized D. None of these
4	Alkaline battery has a voltage of	A. 4.5 V B. 3.5 V C. 2.5 V D. 1.5 V
5	Which one of the following reactions is not spontaneous	
6	Which one of the following reaction takes place spontaneously	
7	Which statement is incorrect about standard hydrogen electrode	A. Its oxidation and reduction potential is zero B. It consists of Pt wire dipped on 1 molar HCl solution C. The electrolyte is 1 molar NaOH solution D. H₂gas is passes in it at 1 atmospheric pressure
8	While balancing an equation by ion electron method, the number of oxygen atoms are balanced by	A. OH⁻ B. H₂O C. O₂ D. H⁺
9	Lead accumulator contains	A. 30% NaCl solution as electrolyte B. 30% HCl solution as electrolyte C. 30% H₂SO₄solution as electrolyte D. 30% NaOH solution as electrolyte
10	In lead accumulator the electrolyte is H_2SO_4solution is	A. 30% H₂SO₄ B. 60% H₂SO₄ C. 80% H₂SO₄ D. 90% H₂SO₄
11	The galvanic or voltaic cells which are rechargeable called as	A. Primary cells B. Secondary cells C. Dry cells D. Infinite cells
12	The voltaic or galvanic cells which cannot be recharged are called	A. Primary cells B. Secondary cells C. Infinite cells D. Fuel cells
13	Electrochemical series is a list of element S arranged into the increasing order of their	A. Standard oxidation potential B. Standard reduction potential C. Cell voltage D. Ionization potential
14	The value of SHE is cathode and anode is always taken to be	A. One B. Zero C. Different D. Same
15	Coupling of Pb with its Pb^{++}/Pb = -0.13 V and Ag with Ag^{+}/ Ag = +0.80 V, the cell reaction	
		A. -0.42 V

16	The reduction potential to copper electrode is +0.34 V and that of Zn electrode is -0.76 V. when these two are coupled the e.m.f. of the cell is	B. +0.42 C. -1.10 V D. +1.10 V
17	A standard hydrogen electrode is used as standard electrode of which electrode potential is arbitrarily taken as	A. +1 B. -1 C. 0.1 D. Zero
18	When a metal is dipped in 1 molar of its solution at 298 K. then potential set up is called	A. Standard electrode potential B. Electric charge C. Ionization potential D. Electroplating
19	The oxidation number of free element is always taken to be	A. 0 B. 1 C. 2 D. -1
20	The function of salt bridge in the galvanic or voltaic cell is to	A. Carry out oxidation at anode B. To carry out reduction at cathode C. Carry out electrolysis D. To prevent the net charge accumulation in either of the half cells
21	The cell which generates electricity as a result of spontaneous oxidation-reduction reaction is called	A. Electrolytic cell B. Nelson's cell C. Galvanic cell D. Down's cell
22	The two half cells of a galvanic cell are connected by	A. Ammeter B. Salt bridge C. Hydrogen electrode D. Copper electrode
23	The process of electrical coating of one metal on another to protect, decorate or to have greater resistance to corrosion is called	A. Electroplating B. Electrolysis C. Conduction D. Induction
24	Purification of an impure copper is made by electrolytic cell, in which impure copper is anode and pure copper is cathode, and the electrolyte used is	A. H₂SO₄ B. CuSO₄ C. ZnSO₄ D. Na₂SO₄
25	Caustic soda is obtained by electrolysis of conc. aqueous solution of NaCl in a cell called	A. Daniell's cell B. Nelson's cell C. Down's cell D. Voltaic cell
26	In the electrolysis of aqueous solution of sodium nitrate, the ions which are reduced at the cathode are	A. H₃O⁺ B. Na⁺ C. OH⁻ D. NO³⁻
27	Metallic conduction is due to the	A. Movement of electrons B. Movement of ions C. Both a and b D. None of these
28	Cell in which an electric current drives a non-spontaneous reaction is called	A. Electrolytic cell B. Voltaic cell C. Biological cell D. Electrochemical cell
29	The electrolysis of $CuSO_4$aqueous solution using copper as cathode as well an anode the substance which deposits at cathode is	A. H₂SO₄ B. Oxygen C. Copper D. Hydrogen
30	By using graphite electrode the electrolysis of aqueous solution of NaCl produces at anode	A. H₂gas B. Cl₂gas C. NaOH D. No metal