ECAT Chemistry Chapter 10 Electrochemistry Online Test

Sr	Questions	Answers Choice
1	Corrosion reaction are	A. Spontaneous redox reaction B. Non-spontaneous acid-base reactions C. Spontaneous acid-base reactions D. None of these
2	Standard reduction of Zn = -0.76 V and that of Ni is -0.25 V. On coupling them by a salt bridge which of these will act as anode	A. Salt bridge will act as anode B. Zn will act as anode C. Ni will act as anode D. None of these
3	Zn does not displace Mg from $MgSO_4$solution because	A. Zn is more electropositive than Mg B. Zn is below Mg in electropositive series C. Zn is above Mg in electrochemical series D. Zn is trivalent Mg is divalent
4	Fe can displace Cu form CuSO4solution because	A. Fe is ferromagnetic B. Fe is below Cu in electrochemical series C. Fe is above Cu in electrochemical series D. Fe exists in divalent oxidation state
5	Which of the following will be good conductor of electricity	A. Pure distilled water B. Molten NaCl C. Dilute solution of glucose D. Chloroform
6	In $KO_2$the oxidation state of oxygen is	A. -2 B. -1 C. +1/2 D. -1/2
7	F_2, Cl_2, $Br_2$and I_2lie below SHE is the Electro chemical series that is why these	A. Undergo reduction B. Undergo oxidation C. Liberate H₂gas with steam D. None of these
8	In the electrolysis of fused bauxite ($Al_2O_32H_2O$) with fused Cryolite (Na_3AlF_6) using carbon rods as anode. The product obtained at cathode is	A. Na metal B. F₂gas C. Al metal D. O₂gas
9	During redox reaction an oxidizing agent	A. Gains electrons B. Is oxidized C. Loses electrons D. Hydrolysed
10	When aqueous solution of NaOH is electrolysed useing graphite electrodes, the product obtained at anode is	A. O₂gas B. H₂gas C. Na metal D. Na₂O
11	In passage of electricity through aqueous solution of $AgNO_3$silver dissolves at anode to form Ag^+, the electrodes are	A. Silver metal B. Pt metal C. Graphite D. Copper metal
12	In the reaction $K_2Cr_2O_7$+ HCl + $CrCl_3$+ Cl_2+ H_2O the element which is reduced is	A. K B. Cl C. Cr D. H
13	Li has the lowest reduction potential while the element with highest reduction potential is	A. H B. F C. O D. N
		A. The reaction is splitted into two half reactions B. H₂O and H<sup>+

14	Which statement is incorrect for balancing of redox reactions by ion-electron method	</sup>ions are added for acidic or neutral reaction to balance O and H atoms C. To balance H, HCi, is added D. To balance O and H in the alkaline reaction OH⁻added
15	Metals like Fe, Mg, Al, Cr, Zn have more negative reduction potentials that is whey	A. These don't react with steam B. These react very slowly with steam to liberate H₂gas C. These react rapidly with steam to produce the metallic oxides and liberate H₂gas D. These react with cold water violently
16	Coinage metals like Au, Pt, Ag and Cu are the least reactive metals and don't liberate H_2gas when treated with acids because	A. These have very high positive values of reduction potentials B. These have very high negative values of reduction potentials C. Their ionization potentials are lowest D. Their reduction potentials are close to SHE
17	Metals which are above SHE in electrochemical series	A. Can liberate H₂from acid B. Cannot liberate H₂from acid C. Cannot always liberate H₂from acid D. None of these
18	Question Image	A. Cu B. H C. N D. O
19	Strong reducing agents gave	A. Greater positive value of standard reduction potential B. Greater negative value of standard reduction potential C. Lesser positive value of standard reaction potential D. None of these
20	In a compound an atom has negative oxidation state because	A. Atom is negatively charged B. Atom acts as cathode C. Atom is more electronegative D. Atom has lowest ionization energy
21	The oxidation number of chromium in $K_2Cr_2O_7$is	A. 14 B. 12 C. 6 D. None of these
22	The oxidation state of an element is zero when	A. It forms an oxide B. It forms hydride C. It is in free state D. Only for noble gases
23	The oxidation number of Mn is $KMnO_4$is	A. +2 B. +4 C. +6 D. +7
24	Oxidation number of oxygen in $OF_2$is	A. +1 B. -1 C. +2 D. -2
25	The oxidation number of H is -1 in the compound	A. H₂O B. H₃BO₃ C. NaOH D. NaH
26	Which statement is correct for the fuel cells	A. These cells operate at low temperature B. These cells operate at low temperature C. No catalyst used for the formation of water D. MnO₂is used as electrolyte
27	Which statement is incorrect for NICAD battery	A. The electrolyte is alkali B. Cd acts as anode C. MnO₂acts as electrolyte

		D. NiO₂acts as cathode
28	Lead accumulator stops discharging current when	A. Lead at anode converted to PbO₂ B. PbO₂at cathode converted to Pb C. Both electrodes are completely covered with PbSO₄ D. Both electrodes are completely covered with Pb(OH)₂
29	Silver oxide battery has a voltage of	A. 2.0 V B. 1.5 V C. 2.5 V D. 1.0 V
30	During electrolysis of KNO_3, $H_2$is evolved	A. Anode B. Cathode C. Both a and b D. None