

## Chemistry Fsc Part 2 Online Test

Sr	Questions	Answers Choice
1	The oxide of beryllium is	A. Acidic B. Basic C. Amphoteric D. None of these
2	Which one of the following does not belong to alkaline earth metals.	A. Be B. Ra C. Ba D. Rn
3	Which oxide is more basic in nature.	A. BeO B. MgO C. CaO D. BaO
4	Which property of hydrogen not resemble to alkali metals.	A. Electronic configuration B. Oxidation state C. Reaction with halogen D. Metallic nature
5	The oxides of metal sare generally	A. Acidic B. Basic C. Neutral D. Amphoteric
6	Which one of the following oxides is amphoteric in nature.	A. MgO B. Na <sub>2</sub> O C. SO <sub>2</sub> D. ZnO
7	Which one of the following elements burns in air to form an oxide which, when shaken with water, give a solution with a pH greater than 7.	A. Carbon B. Magnesium C. sulphur D. Hydrogen
8	Which one of the following elements burns in air to form an oxide which, when shaken with water, give a solution with a pH greater than 7.	A. Carbon B. Magnesium C. sulphur D. Hydrogen
9	Which one of the following elements burns in air to form an oxide which, when shaken with water, give sa solution with a pH greater than 7.	A. Carbon B. Magnesium C. sulphur D. Hydrogen
10	Element of which group reacts with hydrogen and form ionic hydrides.	A. II A B. IV A C. V A D. VI A
11	Which hydride is intermediate in nature.	A. NaH B. BeH <sub>2</sub> C. NH <sub>3</sub> D. HCl
12	Ionic Hydrides react with water to form	A. Proton B. Hydride ions C. Hydroxide ions D. Hydronium ions
13	Which element when react with chlorine form polymeric halide.	A. Na B. Be C. Ba D. P
14	Which has greater hydration energy.	A. Li <sup>+</sup> B. Na <sup>+</sup> C. K <sup>+</sup> D. Mg <sup>+2</sup>
15	In which compound, oxidation state of sulphur is +6	A. H <sub>2</sub> S B. H <sub>2</sub> SO <sub>4</sub> C. H <sub>2</sub> SO <sub>3</sub> D. SO <sub>3</sub>

16	In which group, melting point and boiling point increase downward in a group	A. IA B. II A C. VII A D. Both a and b
17	Variable valency is generally exhibited by	A. Transition elements B. Alkali metals C. s-block elements D. Gaseous elements