

## Chemistry Fsc Part 1 Online Test

Sr	Questions	Answers Choice
1	Colligative properties are used to determine the	A. Freezing points B. Boiling point C. Atomic mass of an element D. Molar mass of solute
2	The vapour pressure of an aqueous solution of glucose is.	A. Equal to vapour pressure of water B. Independent of temperature C. More than vapour pressure of pure water D. Less than vapour pressure of pure water
3	Molal boiling constant for water is 0.52 °C. If 6 g of urea is dissolved in 100 g of water, what will be its boiling point.	A. 100.52 °C B. -100.52 °C C. 100 °C D. 99 °C
4	Molal boiling point elevation depends upon	A. Nature of solvent B. Nature of solute C. Vapour pressure of solution D. None of these
5	The molal boiling point constant is the ratio of elevation of boiling point to	A. Molarity B. Mole fraction of solvent C. Molality D. Mole fraction of solute
6	Solubility of which substance decreases by increasing temperature.	A. NaNO <sub>3</sub> B. KNO <sub>2</sub> C. NaCl D. Ce <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>
7	Solubility of which substance decreases by increasing temperature.	A. NaNO <sub>3</sub> B. KNO <sub>2</sub> C. NaCl D. Ce <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>
8	Solubility curve of CaCl <sub>2</sub> ·6H <sub>2</sub> O shows	A. Decrease in solubility with increase of temperature B. Increase in solubility with increase of temperature C. Discontinuous increase in solubility with temperature D. No effect of temperature on solubility
9	Water and Phenol are partially miscible to each other at room temperature when both liquids are mixed together which is upper layer.	A. Water in Phenol B. Phenol and water C. Pure phenol D. Pure water
10	Which solution is an example of solid in gas	A. Fog B. Steel C. smoke D. Air
11	Butter is solution of	A. Liquid in liquid B. Solid and liquid C. Liquid and solid D. Liquid and gas
12	The temperature which partially immiscible pair of liquid leads to the formation of a single phase is called.	A. Transition temperature B. Absolute temperature C. Consolute temperature D. Room temperature
13	Which pair of mixture is called ideal solution.	A. Chlorobenzene and bromobenzene B. Water alcohol C. Water ether D. HCl and water
		A. High boiling point and high vapour pressure B. High boiling point and low vapour

14	A negative deviation from Raoult's law in solution means, the solution has	pressure C. Low boiling point and low vapour pressure D. Low boiling point and high vapour pressure
15	Azeotropic mixture	A. Obey Raoult's law B. Do not obey Raoult's law C. Boils at low temperature only D. Boils at high temperature only
16	Relative lowering of vapour pressure is equal to.	A. Mole fraction of solute B. Mole fraction of solvent C. Mole fraction of solute and solvent D. Molality of solution
17	In case of non volatile solute, lowering of vapour pressure is proportional to.	A. Mass fraction of solute B. Mole fraction of solvent C. Mole fraction of solute D. None of the above